Determining a Molar Volume

A Safer Experiment



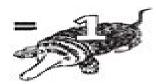
Introduction

The traditional method for calculating molar volumes involves generating oxygen gas—potassium chlorate is heated with numganese dioxide, a catalyst, to produce oxygen gas. This method is quite hazardous because large amounts of oxygen gas are produced and potassium oblorate is a powerful oxiditor of organic materials including the rubber stapper used in the ser-up. In fact, potassium chlorate is a frequent source of accidents on school premises. An easier and rater method presented in this laboratory activity for the calculation of molar volumes involves the use of carbon dioxide instead of oxygen gas.

Concepts

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- · Orlinbay violations
- · Talk all goas have:





Materials

Dry ice, solid carbon disolde, CO₂, small piece about %' square

Balance

Baronseter

Extenuneper flash, Pyrex[®], 125-mL.

Ciral unted cylinder, 500-mL.
Rubber stopper, to fit Erlemmeyer flask.
Thermometer
Timps or insulated gloves (Zetex.*)

Safety Precautions

This activity requires the use of hazardous components and/or has the potential for hazardous reactions. Dry ice can cause freshite; never board dry ice directly with the shin; was appropriate invaluated gloves or use bugs whemever handling dry ice. Do not support the flack until dil of the dry ice has sublimed. Fadure to wast until all old prices are gone may cause the flack to shatter due to increased pressure. When them independ applies graphs, them independ a phone a chemical-resistant appear. Please review current Material Safety Data Shorts for additional safety, handling, and disposal information.

Procedure

- Place the stopper in the Pyres Trienmeyer flask and mass to the nearest 0.01 g. Record this mass in the data table as the mass of the stoppered flask plus air. Recall that the flask is not actually empty—it contains air.
- Remove the stupper. Using tongs, or while wearing insulated gloves, place a small piece of dry ice in the Exhanneyer flask. Do not stupper the flask at this point.
- 3. Gently swirl the flask to aid in the sublimation of the dry ice.
- Immediately after the last bit of day ice has just sublimated and disappeared, rightly stopper the flask. Mass the flask and record this mass in the data table as the mass of the stoppered flask plus CO₂.
- Remove the stopper and fill the Erlemmeyer flack to the top lip with water. Stopper the flack again, Wipe any water off of the outside of the flack.
- Remove the stopper from the flask. Pour the water into a large graduated cylinder and measure the volume of water. The
 volume of water is equal to the volume of air (Step 1) or CO₂ (Step 4) in the flask. Record this volume in liters in the data
 table.
- 7. Record the room temperature in degrees Kelvin and the atmospheric pressure in atm in the data table.
- Record the density of air in the data table. The density of air is both temperature and pressure dependent. This value can be found in a chemistry textbook or reference manual if the temperature and pressure are known.
- 9. Calculate the mass of air in the flask by using the relationship that the density of air (Step 8) is equal to the mass of air

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Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations The American Biology Teacher Nuclear Science Abstracts ,1967 Handbook of Measurement in Science Kentucky Law Journal, 2011 .2005 and Engineering, Volume 3 Myer Kutz, 2016-06-17 A multidisciplinary reference of engineering measurement tools techniques and applications When you can measure what you are speaking about and express it in numbers you know something about it but when you cannot measure it when you cannot express it in numbers your knowledge is of a meager and unsatisfactory kind it may be the beginning of knowledge but you have scarcely in your thoughts advanced to the stage of science Lord Kelvin Measurement is at the heart of any engineering and scientific discipline and job function Whether engineers and scientists are attempting to state requirements quantitatively and demonstrate compliance to track progress and predict results or to analyze costs and benefits they must use the right tools and techniques to produce meaningful data The Handbook of Measurement in Science and Engineering is the most comprehensive up to date reference set on engineering and scientific measurements beyond anything on the market today Encyclopedic in scope Volume 3 covers measurements in physics electrical engineering and chemistry Laser Measurement Techniques Magnetic Force Images using Capacitive Coupling Effect Scanning Tunneling Microscopy Measurement of Light and Color The Detection and Measurement of Ionizing Radiation Measuring Time and Comparing Clocks Laboratory Based Gravity Measurement Cryogenic Measurements Temperature Dependent Fluorescence Measurements Voltage and Current Transducers for Power Systems Electric Power and Energy Measurement Chemometrics for the Engineering and Measurement Sciences Liquid Chromatography Mass Spectroscopy Measurements of Nitrotyrosine Containing Proteins Fluorescence Spectroscopy X Ray Absorption Spectroscopy Nuclear Magnetic Resonance NMR Spectroscopy Near Infrared NIR Spectroscopy Nanomaterials Properties Chemical Sensing Vital for engineers scientists and technical managers in industry and government Handbook of Measurement in Science and Engineering will also prove ideal for academics and researchers at universities and laboratories Manual of Molecular and Clinical Laboratory Immunology John L. Schmitz, Barbara Detrick, Maurice R.

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