

Chemistry Revision

Moles, Empirical formula, Ar and Mr

A_r & M_r

- Relative atomic mass:
 - the weighted average mass of the naturally occurring isotopes of an element as compared to $1/12$ of the mass of carbon-12
 - weighted average means it factors in the abundance of each isotope
 - $(\% \times A_r) + (\% \times A_r) / 100$
- Relative formula mass:
 - the sum of all the atomic masses
 - use relative formula mass to work out percentage composition

Question: find percentage by mass of copper in copper(II) oxide, CuO

A_r : O = 16; Cu = 64

$$\begin{aligned}M_r &= 64 + 16 \\&= 80\end{aligned}$$

$$\begin{aligned}\text{Percentage of copper} &= 64/80 \times 100 \\&= 80\%\end{aligned}$$

Moles

- The number of carbon atoms in exactly 12g of carbon-12
- Formula:
 - number of moles = mass (g) / M_r
 - number of moles = number of particles / 6.02×10^{23}
- Reacting mass calculations:
 - mole (convert mass into moles)
 - mole ratio (from balanced equation)
 - un-mole (convert back to grams)
 - when un-moling in gas calculations, always multiply by 24 instead of M_r

Question: find the mass aluminium oxide when 135g of aluminium is burned with oxygen



- mole = mass/ M_r = $135/27$ = 5 mol Al
- mole ratio = $4:2$ = $2:1$ = 2.5 mol Al_2O_3
- un-mole = 2.5×102 = 255g Al_2O_3

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