

Unit Test 1: Matter, Chemical Trends, & Chemical Bonding**Multiple Choice (10 marks)** (KU: 3, C: 3, I: 2)

Identify the choice that best completes the statement or answers the question. Place your answers in the key below.

1. _____ 2. _____ 3. _____ 4. _____ 5. _____
6. _____ 7. _____ 8. _____ 9. _____ 10. _____

- The element found in the periodic table in Group 6 and Period 5 is
 - tungsten, W
 - molybdenum, Mo
 - tantalum, Ta
 - tellurium, Te
- Which of the following groups is known for having both metallic and non-metallic properties?
 - alkali metals
 - alkali earth metals
 - transition metals
 - metalloids
- How many electrons does a strontium atom have in its outer shell?
 - 1
 - 2
 - 7
 - 8
- Which particles are responsible for the similar chemical and physical properties of elements in the same group?
 - protons and neutrons
 - number of energy shells
 - valence electrons
 - isotopes
- Which of the following chemical formulas most likely represents a pure substance containing ionic bonds?
 - Cl_2
 - SO_2
 - CO_2
 - BaO
- Consider the equation $\text{X}_{\text{g}} + e^- \rightarrow \text{X}^- + \text{energy}$. The "energy" term in the equation represents
 - electronegativity
 - electron affinity
 - ionization energy
 - nuclear charge
- Elements A, B, C, and D (found in Groups 1–17) have atomic radii of $265 \mu\text{m}$, $160 \mu\text{m}$, $185 \mu\text{m}$, and $175 \mu\text{m}$, respectively. Which element will most likely have the lowest ionization energy?
 - A
 - B
 - C
 - D
- List the following elements from smallest to largest atomic size:
I. barium II. strontium III. bismuth IV. iodine
 - III, I, II, IV
 - IV, III, II, I
 - IV, II, III, I
 - IV, II, I, III
- Which shape and bond angle is predicted by VSEPR theory for NF_3 ?
 - trigonal pyramid, less than 109.5°
 - tetrahedral, 109.5°
 - trigonal planar, 109.5°
 - trigonal planar, 120°

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Table of Contents Grade 1chemistry Unit Practice Test

1. Understanding the eBook Grade 1chemistry Unit Practice Test
 - The Rise of Digital Reading Grade 1chemistry Unit Practice Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 1chemistry Unit Practice Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 1chemistry Unit Practice Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 1chemistry Unit Practice Test

- Personalized Recommendations
- Grade 1chemistry Unit Practice Test User Reviews and Ratings
- Grade 1chemistry Unit Practice Test and Bestseller Lists
- 5. Accessing Grade 1chemistry Unit Practice Test Free and Paid eBooks
 - Grade 1chemistry Unit Practice Test Public Domain eBooks
 - Grade 1chemistry Unit Practice Test eBook Subscription Services
 - Grade 1chemistry Unit Practice Test Budget-Friendly Options
- 6. Navigating Grade 1chemistry Unit Practice Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Grade 1chemistry Unit Practice Test Compatibility with Devices
 - Grade 1chemistry Unit Practice Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 1chemistry Unit Practice Test
 - Highlighting and Note-Taking Grade 1chemistry Unit Practice Test
 - Interactive Elements Grade 1chemistry Unit Practice Test
- 8. Staying Engaged with Grade 1chemistry Unit Practice Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 1chemistry Unit Practice Test
- 9. Balancing eBooks and Physical Books Grade 1chemistry Unit Practice Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 1chemistry Unit Practice Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Grade 1chemistry Unit Practice Test
 - Setting Reading Goals Grade 1chemistry Unit Practice Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade 1chemistry Unit Practice Test

- Fact-Checking eBook Content of Grade 1chemistry Unit Practice Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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