Chapter 6 - Lecture Worksheet 3

Using standard heats of formation, calculate the standard enthalpy change for the following reaction.

$$CO_2(g) + 4H_2(g)$$
 ----> $CH_4(g) + 2H_2O(g)$

$$\Delta H^{\circ}_{r=} = \Delta H^{\circ}_{f}(CH_{4},g) + 2\Delta H^{\circ}_{f}(H_{2}O,g) - \Delta H^{\circ}_{f}(CO_{2},g) - 4\Delta H^{\circ}_{f}(H_{2},g)$$
$$= -75 + 2(-242) - (-394) - 4(0) \text{ kJ} = -165 \text{ kJ}$$

(1) -711 kJ (2) -165 kJ (3) -15 kJ (4) 15 kJ (5) 165 kJ (6) 229 kJ (7) 711 kJ

Using average bond enthalpies estimate the standard enthalpy change for the same reaction.

$$CO_2(g) + 4H_2(g)$$
 ----> $CH_4(g) + 2H_2O(g)$

$$\Delta H^0_{\text{res}} = 2BE(C = O) + 4BE(H - H) - 4BE(C - H) - 4BE(O - H)$$

= 2(728) + 4(436) - 4(413) - 4(463) kJ = -304 kJ

Which is more accurate? Why?

AH° calculated from ΔH°_f is more accurate because ΔH° calculated from bond enthalpies uses AVERAGE bond enthalpies obtained from a number of different compounds in which the bonds arise and are not specific to the compound in question.

	StandardEnthalpy of formation (kJ/mol)
CHL(g)	-75
CO ₂ (g)	-394
$C_pH_p(g)$	+227
$H_{if}O(g)$	-242
H ₂ O(I)	-286
Nittedge)	-46
HFC3(g)	-92
HICh(mq)	-167
CarCO ₂ (x)	-1208
Car(Offly(sep)	-1003
Car(OHD)(s)	-986
CiarO(s)	-635
SO ₂ (g)	-297
S(O _v (g))	-396

	Average Bond Enthalpies (kJ/mol
316-316	436
284-88	391
260 - 260	161
250-050	418
NeN	945
O-H	463
C-O	3.51
CHO	728
CwO	1072
CI-CI	242
CI-HI	432
Bir-Bir	193
Be-H	366
C-H	413

Enthalpy Worksheet With Answers

Anthony S. Fauci



Decoding Enthalpy Worksheet With Answers: Revealing the Captivating Potential of Verbal Expression

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