


Surajit Das · Hirak Ranjan Dash

# Microbial Biotechnology– A Laboratory Manual for Bacterial Systems

 Springer

# Microbial Biotechnology A Laboratory Manual For Bacterial Systems

**IM Harris**



## **Microbial Biotechnology A Laboratory Manual For Bacterial Systems:**

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 $2 \text{ CO}_2 \text{ g} + \text{H}_2\text{O g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + \frac{5}{2} \text{O}_2 \text{ g}$   $\Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O g} + \text{H}_2 \text{ g} \rightarrow \frac{1}{2} \text{O}_2 \text{ g}$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow \text{CO}_2 \text{ g} + 3 \text{H}_2\text{O g}$   $\Delta h = 283 \text{ kJ}$

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web chemistry 120 hess s law worksheet calculate  $\Delta h$  for the reaction  $\text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$  from the following data  
 $\text{C}_2\text{H}_4 \text{ g} + 3 \text{O}_2 \text{ g} \rightarrow 2 \text{CO}_2 \text{ g} + 2 \text{H}_2\text{O l}$   $\Delta h = 1411 \text{ kJ}$  mole  
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow 2 \text{CO}_2 \text{ g} + 3 \text{H}_2\text{O l}$   $\Delta h = 1560 \text{ kJ}$  mole  
 $\text{H}_2 \text{ g} + \frac{1}{2} \text{O}_2 \text{ g} \rightarrow \text{H}_2\text{O l}$   $\Delta h = 285 \text{ kJ}$  mole

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hess s law example 12.1 determine  $\Delta h$  for the target reaction  
 $\text{C}_2\text{H}_4 \text{ g} + 3 \text{O}_2 \text{ g} \rightarrow 2 \text{CO}_2 \text{ g} + 2 \text{H}_2\text{O g}$   
given the following information

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hess s law is written as  $\Delta h_{\text{target}} = \Delta h_1 + \Delta h_2 + \Delta h_3$  or  $\Delta h_{\text{target}} = \sum \Delta h_{\text{known}}$

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to use hess s law two principles must be understood  
one if an equation is reversed the sign of the  $\Delta h$  value is also reversed  
two if an equation is multiplied by a

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 $\text{Mg s} + \frac{1}{2} \text{O}_2 \text{ g} \rightarrow \text{MgO s}$  we used

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cohen 2016 the equations for the reactions in the experiment done are as follows  
 $\text{NaOH s} + \text{Na aq} \rightarrow \text{OH aq}$

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web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta h_{\text{products}}$

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 $\text{1 NaOH (s)} + \text{Na}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$   
 $\text{2 NaOH (s)} + \text{H}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$   
 $\text{3 NaOH (s)} + \text{H}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$

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