

More Practice with Gas Laws **KEY**

Directions: For each question, identify the applicable law and solve for the correct answer using dimensional analysis. Express your answer to the correct number of significant figures. Use scientific notation where appropriate. Record all work in your science notebooks.

I. Gas Law Problems

1. A sample of hydrogen gas occupies a volume of 1.8 L at 0.80 atm. What would its volume be at 1.0 atm?

Boyle's Law $P_1V_1 = P_2V_2$

$P_1 = 0.80 \text{ atm}$ $V_1 = 1.8 \text{ L}$

$P_2 = 1.0 \text{ atm}$ $V_2 = ??$

$(.8 \text{ atm})(1.8 \text{ L}) = (1.0 \text{ atm})(V_2)$

$1.44 \text{ atm} \times \text{L} = 1 \text{ atm } V_2$

$V_2 = 1.4 \text{ L}$

2. A balloon full of helium has a volume of 1.75 L at a temperature of 28°C. What would the balloon's volume be at 50°C?

Charles' Law

$\frac{V_1}{T_1} = \frac{V_2}{T_2}$

$V_1 = 1.75 \text{ L}$ $T_1 = 301 \text{ K}$

$V_2 = ??$ $T_2 = 323 \text{ K}$

$V_2 = 1.9 \text{ L}$

3. If you have 5.00 mol of fluorine gas at 273 K and 1.00 atm of pressure, what will be the volume of the gas?

Ideal Gas Law

$PV = nRT$

$(1.00 \text{ atm})(V) = 5.00 \text{ mol } (.0821 \text{ L} \times \text{atm/mol} \times \text{K})(273 \text{ K})$

$V = 112 \text{ L} = 1.12 \times 10^2 \text{ L}$

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AW Chickering



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