

## Chemistry Periodic Table Test Chapter 6 Chem IH

**Matching** Match each item with the correct statement below.

- |                      |                     |
|----------------------|---------------------|
| a. electronegativity | e. transition metal |
| b. ionization energy | f. periodic law     |
| c. electrons         | g. group            |
| d. metal             | h. period           |

- \_\_\_\_ 1. vertical column in the periodic table
- \_\_\_\_ 2. type of element characterized by the presence of electrons in the *d* orbital
- \_\_\_\_ 3. type of element characterized by losing electrons easily
- \_\_\_\_ 4. ability of an atom to attract electrons when the atom is in a compound
- \_\_\_\_ 5. energy required to remove an electron from an atom

### Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- \_\_\_\_ 6. What is another name for the main group elements?
- |                              |                         |
|------------------------------|-------------------------|
| a. s block elements          | c. s & p block elements |
| b. p block elements elements | d. d block elements     |
- \_\_\_\_ 7. What is another name for the transition metals?
- |                              |                              |
|------------------------------|------------------------------|
| a. noble gases               | c. p block elements elements |
| b. s block elements elements | d. d block elements          |
- \_\_\_\_ 8. Who arranged the elements according to atomic mass and used the arrangement to predict the properties of missing elements?
- |                      |                     |
|----------------------|---------------------|
| a. Henry Moseley     | c. John Dalton      |
| b. Antoine Lavoisier | d. Dmitri Mendeleev |
- \_\_\_\_ 9. Which of the following categories includes the majority of the elements?
- |               |              |
|---------------|--------------|
| a. metalloids | c. metals    |
| b. liquids    | d. nonmetals |
- \_\_\_\_ 10. In which of the following sets is the symbol of the element, the number of protons, and the number of electrons given correctly?
- |                                 |                                    |
|---------------------------------|------------------------------------|
| a. In, 49 protons, 49 electrons | c. Cs, 55 protons, 132.9 electrons |
| b. Zn, 30 protons, 60 electrons | d. F, 19 protons, 19 electrons     |
- \_\_\_\_ 11. The atomic number of an element is the total number of which particles in the nucleus?
- |             |                          |
|-------------|--------------------------|
| a. neutrons | c. electrons             |
| b. protons  | d. protons and electrons |
- \_\_\_\_ 12. What element has the electron configuration  $1s^2 2s^2 2p^6 3s^2 3p^2$ ?
- |             |            |
|-------------|------------|
| a. nitrogen | c. silicon |
| b. selenium | d. silver  |
- \_\_\_\_ 13. Which subatomic particle plays the greatest part in determining the properties of an element?
- |             |                      |
|-------------|----------------------|
| a. proton   | c. neutron           |
| b. electron | d. none of the above |
- \_\_\_\_ 14. Which of the following elements is a transition metal?
- |           |              |
|-----------|--------------|
| a. cesium | c. tellurium |
| b. copper | d. tin       |

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appeal to experts and interested lay persons alike The Periodic Table begins with an overview of the importance of the periodic table and of the elements and it examines the manner in which the term element has been interpreted by chemists and philosophers The book then turns to a systematic account of the early developments that led to the classification of the elements including the work of Lavoisier Boyle and Dalton and Cannizzaro The precursors to the periodic system like D bereiner and Gmelin are discussed In chapter 3 the discovery of the periodic system by six independent scientists is examined in detail Two chapters are devoted to the discoveries of Mendeleev the leading discoverer including his predictions of new elements and his accommodation of already existing elements Chapters 6 and 7 consider the impact of physics including the discoveries of radioactivity and isotopy and successive theories of the electron including Bohr s quantum theoretical approach Chapter 8 discusses the response to the new physical theories by chemists such as Lewis and Bury who were able to draw on detailed chemical knowledge to correct some of the early electronic configurations published by Bohr and others Chapter 9 provides a critical analysis of the extent to which modern quantum mechanics is or is not able to explain the periodic system from first principles Finally chapter 10 considers the way that the elements evolved following the Big Bang and in the interior of stars The book closes with an examination of further chemical aspects including lesser known trends within the periodic system such as the knight s move relationship and secondary periodicity as well as attempts to explain such trends

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