CHEM 101 Worksheet, Chapter 6: Thermochemistry

Multiple choice questions 1-10 (page 1): circle the correct answer(A-E)

Show your work questions 11-13 (pages 2-3): give the complete answer in the space provided R = 8.314 J/(mol.K) = 0.08206 L.atm(mol.K) 1 cal = 4.184 J 0 °C = 273 K

Calculate the heat effect q (in kJ) when 112 g of iron (Fe(s)) is heated from 20 °C to 300 °C. The specific heat of Fe(s) = 0.48 J/g.K

A. -34.9 kJ kJ

B. +34.9 kJ

 $C_{\rm s} = 15.0 \; {\rm kJ}$

D. -270 J

E. +15.0 kJ

 The dissolution reaction for ammonium chloride: NH₄Cl(s) → NH₄'(aq) + Cl⁻(aq) is endothermic. When ammonium chloride is dissolved in water in an insulated (adiabatic) system, the temperature will:

A. increase

B. decrease

C. stay the same

3. Consider the reaction: $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$. $\Delta H_{rm} = +178 \text{ kJ}$ What is the enthalpy change when 20.0 kg CaO is formed? A. $+6.36\times10^s \text{ kJ}$ B. $-6.36\times10^s \text{ kJ}$ C. $+6.36\times10^s \text{ J}$ D. 63.6 kJ E. +178 kJ4. Consider the reaction: $2NO + O_2 \rightarrow 2NO_2$ $\Delta H_{rm} = -114.6 \text{ kJ}$ As always, the ΔH is "for the reaction as written"! Calculate the enthalpy change ΔH , when 12.6 g NO_2 is produced. A. -31.4 kJ B. +31.4 kJ C. -15.7 kJ D. +15.7 kJ E. none of these

5. How much heat is associated with the reaction that produces 256.1 g of ammonia, NH₃? $N_2(g) + 3 H_2(g) \rightarrow 2 NH_2(g) = \Delta H^2_{cm} = -92.45 \text{ kJ}$

A. 92.45 kJ of heat is released B. 696.1 kJ of heat is released C. 184.9 kJ of heat is absorbed D. 184.9 kJ of heat is released E. 92 kJ of heat is released

6. Which two of the following are not formation reactions?

1. $\frac{1}{2}N_{2}(g) + \frac{1}{2}H_{2}(g) \rightarrow NH_{3}(g)$ 2. $C(s) + O_{2}(g) \rightarrow CO_{2}(g)$ 3. $2SO_{2}(g) + O_{2}(g) \rightarrow 2SO_{3}(g)$ 4. $2Fe(s) + \frac{1}{2}O_{2}(g) \rightarrow Fe_{2}O_{3}(s)$ 5. $\frac{1}{2}H_{2}(g) + \frac{1}{2}I_{2}(g) \rightarrow HI(g)$ A. 1 and 2 B. 3 and 4 C. 2 and 3 D. 4 and 5 E. 3 and 5 7. Consider the thermochemical reaction equation: $2CO(g) + O_{2}(g) \rightarrow 2CO_{2}(g) \rightarrow H = -198 \text{ kJ}$ What is the enthalpy change for the reaction: $CO_{2}(g) \rightarrow CO(g) + \frac{1}{2}O_{2}(g)$ A. +198 kJ B. -99 kJ C. +99 kJ D. +396 kJ E. -198 kJ

8. Calculate the enthalpy change for the reaction: $2KClO_3 \rightarrow 2\ KCl + O_2(g)$ Standard enthalpy of formation values: $\Delta H_i^*(KClO_3) = -389\ kJ/mol\ \Delta H_i^*(KCl) = -436\ kJ/mol\ A. -94\ kJ$ B. -47 kJ C. +94 kJ D. +47 kJ

E. Cannot be calculated because ΔH_f°(O₂(g)) is not given

9. Given the following two thermochemical equations:

E. -668 kJ

10. Which of the following statements is correct for a combustion reaction in an adiabatic "bomb" calorimeter?

A. the pressure is constant B. the temperature is constant C the volume is constant D. the calorimeter constant's C units are J/g. E. $\Delta T(K) = \Delta T(^{\circ}C) + 273.15$ Show your work for the following questions.

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Table of Contents Multiple Choice Thermochemistry Questions

- 1. Understanding the eBook Multiple Choice Thermochemistry Questions
 - The Rise of Digital Reading Multiple Choice Thermochemistry Questions
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Multiple Choice Thermochemistry Questions
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Thermochemistry Questions
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Multiple Choice Thermochemistry Questions
 - Personalized Recommendations
 - Multiple Choice Thermochemistry Questions User Reviews and Ratings
 - Multiple Choice Thermochemistry Questions and Bestseller Lists
- 5. Accessing Multiple Choice Thermochemistry Questions Free and Paid eBooks
 - Multiple Choice Thermochemistry Questions Public Domain eBooks
 - Multiple Choice Thermochemistry Questions eBook Subscription Services
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 - Multiple Choice Thermochemistry Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Thermochemistry Questions
 - Highlighting and Note-Taking Multiple Choice Thermochemistry Questions
 - Interactive Elements Multiple Choice Thermochemistry Questions
- 8. Staying Engaged with Multiple Choice Thermochemistry Questions

- o Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Multiple Choice Thermochemistry Questions
- 9. Balancing eBooks and Physical Books Multiple Choice Thermochemistry Questions
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Multiple Choice Thermochemistry Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Multiple Choice Thermochemistry Questions
 - Setting Reading Goals Multiple Choice Thermochemistry Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Multiple Choice Thermochemistry Questions
 - Fact-Checking eBook Content of Multiple Choice Thermochemistry Questions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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