

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.



- **Second Law of Thermodynamics**: The entropy of an isolated system never decreases.
 - **Entropy** is a measure of the disorder or randomness of a system.

$$\Delta S = \frac{Q}{T}$$

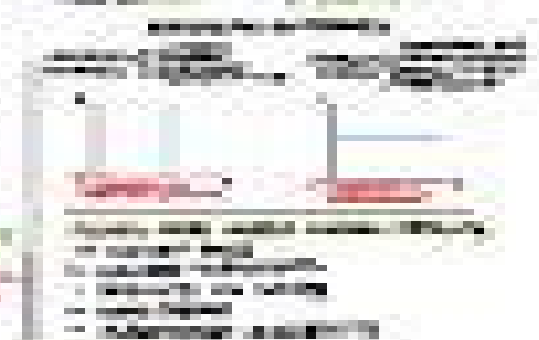
 where ΔS is the change in entropy, Q is the heat added, and T is the temperature.



- **Third Law of Thermodynamics**: The entropy of a perfect crystal at absolute zero is zero.
 - **Heat Engines** convert heat into work.

$$\eta = \frac{W}{Q}$$

 where η is the efficiency, W is the work done, and Q is the heat added.



Microeconomics Study Guide Final

Rachel S Tattersall



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