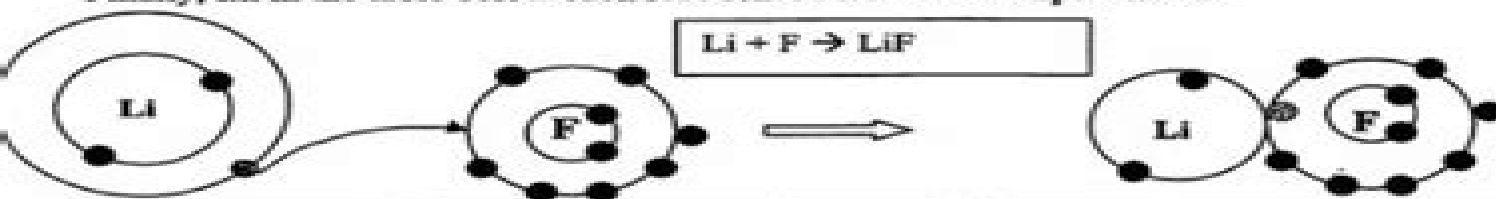


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For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1	1 to F	Li^+
F	7	1 from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$	Li	1	1 to Cl	Li^+
	Cl	7	1 from Li	Cl^-
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$	Ca	2	2 to O	Ca^{2+}
	O	6	2 from Ca	O^{2-}
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$	Be	2	1 to each F	Be^{2+}
	F	7	Each 1 gets 1 from Be	F^-
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$	Mg	2	2 to S	Mg^{2+}
	S	6	2 from Mg	S^{2-}
5) $\text{K} + \text{F} \Rightarrow \text{KF}$	K	1	1 to F	K^+
	F	7	1 from K	F^-

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N Colangelo

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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