energy level

CHAPTER 7 STUDY GUIDE

Ionic Compounds and Metals

Section 7.1 Ion Formation

chemical bond

In your textbook, read about chemical bonds and formation of ions.

electrons

Use each of the terms below just once to complete the passage.

54770054C5C5C			D. 2002 - 100 - 100
nucleus	octet	pseudo-noble gas formations	valence
The force t	that holds two atoms together i	s called a(n) (1)	
Such an attacl	ament may form by the attracti	on of the positively charged	
(2)	of one atom for the negatively charged		
(3)	of another atom, or by the attraction of charged atoms, which are called		
(4)	The attractions may also involve		
(5)	electrons, which are the electrons in the outermost		
(6)	The (7)	are a family of elements	that have very
little tendency	to react. Most of these elemen	its have a set of eight outermost electrons	, which is called a
stable (8)	. The rel	atively stable electron structures develope	ed by loss of
electrons in ex	ertain elements of groups 3, 4,	13, and 14 are called (9)	
For each stat	ement below, write true or fa	lse.	
	 A positively charged ion is called an anion. 		
	 Elements in group 1 le charge. 	se their one valence electron, forming an	ion with a 1+
	_ 12. Elements tend to react	so that they acquire the electron structure	e of a halogen.
	_ 13. A sodium atom tends t	o lose one electron when it reacts.	
	14. The electron structure formation.	of a zinc ion (Zn2*) is an example of a ps	eudo- noble gas
	_ 15. A Cl* ion is an example	e of a cation.	
	_ 16. The ending -ide is use	d to designate an anion.	
	 Nonmetals form a stab becoming anions. 	de outer electron configuration by losing	electrons and

noble gases

pouts.

Ionic Compounds And Metals Study Guide

Dr Rajesh Thakur, Dr S R Singh, Subhash Jain, Mamta Mehrotra

Ionic Compounds And Metals Study Guide:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of guestions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students A-Level Study Guide Chemistry Ed H2.2 CS Toh, 2016-03-08 This is an ebook version of the A Level Study Guide Chemistry Higher 2 Ed H2 2 published by Step by Step International Pte Ltd For the revised Higher 2 H2 syllabus with first exam in 2017 This ebook gives concise illustrated notes and worked examples It is intended as a study guide for readers who have studied the O Level Chemistry or the equivalent It contains material that most readers should want to take note of when attending formal lessons and or discussions on the Singapore Cambridge GCE A Level Higher 2 H2 Chemistry As the Higher 1 H1 Chemistry syllabus is a subset of the H2 Chemistry syllabus this ebook is also suitable for readers studying Chemistry at the H1 level The concise notes cover

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