

# IONIC

## DEFINITION

an electrical attraction between a positively charged ion  $\oplus$  and a negatively charged ion  $\ominus$



more protons  $\oplus$  than electrons  $\ominus$

atom loses electron to become  $\oplus$  charge

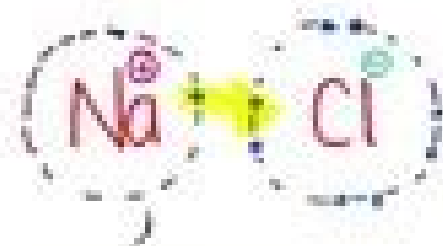


more electrons  $\ominus$  than protons  $\oplus$

atom gains electrons to become  $\ominus$  charge

## PROPERTIES

- high boiling point
- high melting point
- solid @ room temperature
- dissolve in water
- conduct electricity
- brittle



# COVALENT

## DEFINITION

chemical bond formed when electrons are shared.

single  $\longleftrightarrow$  share one pair of electrons

double  $\longleftrightarrow$  share two pairs of electrons

triple  $\longleftrightarrow$  shares three pairs of electrons

## PROPERTIES

- do not conduct electricity
- lower melting point
- do not separate in water
- lower boiling point
- can be solid, liquid or gas at room temperature

shared electrons circle both atoms



# METALLIC

## DEFINITION

bond formed when many metal atoms share their pooled electrons

LIKE IONIC

give away electrons

LIKE COVALENT

metals share electrons

## PROPERTIES

- high boiling point
- high melting point
- good conductors of heat & electricity
- shiny



# Ionic Bonding Notes Guide

**Camilla Rothe**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Ionic Bonding Notes Guide:**

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