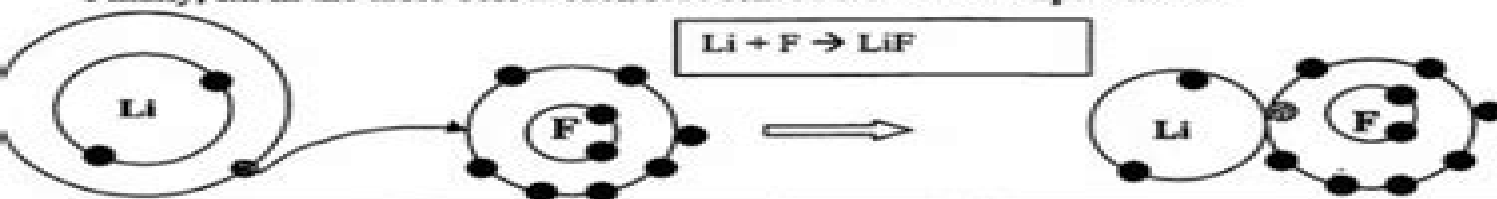


# Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1	1 to $\text{F}$	$\text{Li}^+$
F	7	1 from Li	$\text{F}^-$

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$	Li	1	1 to Cl	$\text{Li}^+$
	Cl	7	1 from Li	$\text{Cl}^-$
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$	Ca	2	2 to O	$\text{Ca}^{2+}$
	O	6	2 from Ca	$\text{O}^{2-}$
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$	Be	2	1 to each F	$\text{Be}^{2+}$
	F	7	Each 1 gets 1 from Be	$\text{F}^-$
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$	Mg	2	2 to S	$\text{Mg}^{2+}$
	S	6	2 from Mg	$\text{S}^{2-}$
5) $\text{K} + \text{F} \Rightarrow \text{KF}$	K	1	1 to F	$\text{K}^+$
	F	7	1 from K	$\text{F}^-$

# Ionic Bonds Section 7 2 Worksheet Answers

**Samuel N. Namowitz, Nancy E.  
Spaulding**



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