#### Chapter 7: Ionic Compounds and Metals

#### 7.1 Ion Formation p. 206

- A chemical bond is the force that holds two atoms together.
- Chemical bonds form by the attraction between the positive nucleus of one atom and the negative electrons of another atom.
- Atom's try to form the <u>octet</u>—the stable arrangement of eight valence electrons in the outer energy level—by gaining or losing valence electrons.
- 4. A positively charged ion is called a cation.
- 5. An example is when sodium atom loses one valence electron to become a cation ion.
- Metals are reactive because they lose valence electrons easily.
- 7. Transition metals commonly form 2+ or 3+ ions, but can form greater than 3+ ions.
- Other relatively stable electron arrangements are referred to as <u>pseudo-noble gas</u> configurations.
- 9. An anion is a negatively charged ion.
- An example would be chlorine gaining an electron to become a chlorine ion.
- 11. Nonmetal ions gain the number of electrons required to fill an octet.
- Some nonmetals can gain or lose electrons to complete an octet.

#### 7.2 Ionic Bonds and Ionic Compounds p.210

- The electrostatic force that holds oppositely charged particles together in an ionic compound is called an ionic bond.
- 2. Compounds that contain ionic bonds are called ionic compounds.
- Binary ionic compounds contain only two different elements—a metallic cation and a nonmetallic anion.
- Formation of an ionic bond (table 7.4 p.211)
- Positive and negative ions exist in a ratio determined by the number of electrons transferred from the metal atom to the non-metal atom.
- The repeating pattern of particle packing in an ionic compound is called an ionic crystal.
- The strong attractions among the positive and negative ions result in the formation of the crystal lattice.
- A <u>crystal lattice</u> is the three-dimensional geometric arrangement of particles, and is responsible for the structure of many minerals.
- 9. Melting point, boiling point, and hardness depend on the strength of the attraction.
- 10. Melting and boiling points of some ionic compounds table 7.5 p.215.
- In a <u>solid, ions are locked</u> into position and electrons <u>cannot flow freely</u>—solid ions are <u>poor</u> conductors of electricity.
- Liquid ions or ions in aqueous solution have electrons that are <u>free to move</u>, so they <u>conduct</u> electricity easily.
- An ion in agueous solution that conducts electricity is an electrolyte.
- Reactions that absorb energy are endothermic.
- Reactions that release energy are exothermic.
- The energy required to separate 1 mol of ions in an ionic compound is referred to as the <u>lattice</u> energy.
- 17. Lattice energy is directly related to the size of the ions that are bonded.

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