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SET-2

प्रश्न-पत्र कोड Q.P. Code 58/5/2

रोल नं. Roll No.

परीक्षार्थी प्रश्न-पत्र कोड को उत्तर-पुस्तिका के म्ख-पृष्ठ पर अवश्य लिखें।

Candidates must write the Q.P. Code on the title page of the answer-book.

अर्थशास्त्र (सैद्धान्तिक)

ECONOMICS (Theory)

निर्धारित समय: ३ घण्टे

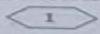
Maximum Marks: 80

Time allowed: 3 hours

- कपया जाँच कर लें कि इस प्रश्न-पत्र में मुद्रित पृष्ठ 23 हैं।
- प्रश्न-पत्र में दाहिने हाथ की ओर दिए गए प्रश्न-पत्र कोड को परीक्षार्थी उत्तर-पस्तिका के मुख-पुष्ठ पर लिखें।
- कृपया जाँच कर लें कि इस प्रश्न-पत्र में 34 प्रश्न हैं।
- कृपवा प्रश्न का उत्तर लिखना शुरू करने से पहले, उत्तर-पुस्तिका में प्रश्न का क्रमांक अवश्य लिखें।
- इस प्रश्न-पत्र को पढ़ने के लिए 15 मिनट का समय दिया गया है। प्रश्न-पत्र का वितरण पूर्वाह में 10.15 बजे किया जाएगा । 10:15 बजे से 10:30 बजे तक परीक्षार्थी केवल प्रश्न-पत्र को पहेंगे और इस अबधि के दौरान वे उत्तर-पुस्तिका पर कोई उत्तर नहीं लिखेंगे।
- Please check that this question paper contains 23 printed pages.
- Q.P. Code given on the right hand side of the question paper should be written on the title page of the answer-book by the candidate.
- Please check that this question paper contains 34 questions.
- Please write down the serial number of the question in the answerbook before attempting it.
- 15 minute time has been allotted to read this question paper. The question paper will be distributed at 10.15 a.m. From 10.15 a.m. to 10.30 a.m., the candidates will read the question paper only and will not write any answer on the answer-book during this period.



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Michael Hoelscher

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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of temperature rise that being 1 267 per degree in 1874 victor regnault obtained better experimental results showing this number to be 1 273 and concluded that the pressure would approach zero at 1 273 degrees c

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web the volume of 1 00 mol 1 00 mol of any gas at stp standard temperature 273 15 k and pressure 1 atm is measured to be 22 414l 22 414 l we can substitute 101 325kpa 101 325 kpa for pressure 22 414 l 22 414 l for volume and 273 15 k 273 15 k for temperature into the ideal gas equation and solve for r r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of gases $273\ 15\ k$ and $1\ atm\ 101\ 325\ kpa\ 1$ at stp one mole of an ideal gas has a volume of about $22\ 4\ l$ this is referred to as the standard molar volume figure $9\ 18$

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