

## POGIL: Periodic Table Trends

(Adapted from Rush Henrietta CSD)

### Why:

The Periodic Table is one of the greatest inventions in the history of man. It allows scientists to predict physical and chemical properties of the elements. Dimitri Mendeleev (a Russian scientist) and Robert Mosley (a British chemist) put together this table in the late 1800's based on properties of the elements known at the time. The trends in these properties as you go across periods and down groups is the subject of this Chem POGIL.

Periodic Table of the Elements

Legend:

- s-block
- p-block
- d-block
- f-block
- noble gases

### Success Criteria:

- Understand the meaning of *atomic radius, reactivity, electronegativity, ionization energy*
- Recognize trends in atomic radius, reactivity, ionization energy and electronegativity as you go across periods and down groups.

### Resources:

Reference Table S, Periodic Table, Unit 3 (Chapter 6) Notes, & Chapter 6 in textbook

### Prerequisites:

Answer the following questions using the resources above.

- What happens to the number of valence electrons as you go down a group on the periodic table?
- What happens to the number of valence electrons as you go from left to right across a period?
- What happens to the number of energy levels as you go down a group on the periodic table?
- What happens to the number of energy levels as you go from left to right across a period?
- Draw a Lewis structure for a Cl atom.
- Draw a Bohr diagram for the Cl atom. (Draw above!)
- In this lab we will be talking about atomic radius.
  - Look up the atomic radius of Cl on Table S: \_\_\_\_\_
  - Why is it possible to show the atomic radius using a Bohr diagram but not by using a Lewis structure?
  - On your Bohr diagram draw an arrow to represent the radius of the Cl atom.
- Define *periodic*:
- Define *trend*:

# Geometry Connections Volume 2 Answers

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## **Table of Contents Geometry Connections Volume 2 Answers**

1. Understanding the eBook Geometry Connections Volume 2 Answers
  - The Rise of Digital Reading Geometry Connections Volume 2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Geometry Connections Volume 2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Geometry Connections Volume 2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Geometry Connections Volume 2 Answers
  - Personalized Recommendations
  - Geometry Connections Volume 2 Answers User Reviews and Ratings
  - Geometry Connections Volume 2 Answers and Bestseller Lists
5. Accessing Geometry Connections Volume 2 Answers Free and Paid eBooks
  - Geometry Connections Volume 2 Answers Public Domain eBooks
  - Geometry Connections Volume 2 Answers eBook Subscription Services
  - Geometry Connections Volume 2 Answers Budget-Friendly Options
6. Navigating Geometry Connections Volume 2 Answers eBook Formats

- ePub, PDF, MOBI, and More
- Geometry Connections Volume 2 Answers Compatibility with Devices
- Geometry Connections Volume 2 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Geometry Connections Volume 2 Answers
  - Highlighting and Note-Taking Geometry Connections Volume 2 Answers
  - Interactive Elements Geometry Connections Volume 2 Answers
- 8. Staying Engaged with Geometry Connections Volume 2 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Geometry Connections Volume 2 Answers
- 9. Balancing eBooks and Physical Books Geometry Connections Volume 2 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Geometry Connections Volume 2 Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Geometry Connections Volume 2 Answers
  - Setting Reading Goals Geometry Connections Volume 2 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Geometry Connections Volume 2 Answers
  - Fact-Checking eBook Content of Geometry Connections Volume 2 Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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