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web an solution initially contains a catalytic amount of an enzyme with k m 1 5 mm 0 25 m of substrate and no product after



45 seconds the solution contains 25  $\mu\text{m}$  of product find  $v_{\text{max}}$  and the concentration of product after 20 minutes hint  $s/k/m$   
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web a reaction and its experimental rate law are represented below  $2\text{ICl} + \text{H}_2 \rightarrow 2\text{HCl} + \text{I}_2$  rate  $k[\text{ICl}]^2$  a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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web solution a first order with respect to  $\text{Br}_2$  first order with respect to  $\text{BrO}_3^-$  and second order with respect to  $\text{H}^+$  hence the overall order of the reaction is equal to 1 + 1 + 2 = 4 b order of the reaction with respect to acetaldehyde is 3/2 and overall order is also 3/2

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web dec 26 2015 in the nuclear industry workers use a rule of thumb that the radioactivity from any sample will be relatively harmless after 10 half lives calculate the fraction of a radioactive sample that remains after this time period hint radioactive decay obeys first

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web in order to have successful collisions the colliding particles must have both the proper amount of energy and the proper fill in blank draw an energy diagram which shows the shape of the curve you might expect for the reaction in this question the overall reaction is exothermic

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web mar 13 2023 work in groups on these problems you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 for the reaction below if substance a is disappearing at a rate of 1.82 times  $10^{-2} \text{ mol l}^{-1} \text{ s}^{-1}$  at what rate is c appearing 3 a 3 b  $\rightarrow$  5 c 2 d

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web reaction rates google classroom mg s 2 hcl a q mgcl a 2 a q h a 2 g the reaction between mg s and hcl a q is represented by the equation above in a kinetics experiment a 0.080 g sample of mg s ribbon is added to an erlenmeyer flask containing 50 ml of 3.0 M hcl a q at 25 °C

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