

How Substances Dissolve

KEY IDEAS

As you read this section, keep these questions in mind:

- Why is water called the universal solvent?
- How do substances dissolve?

Why Do Substances Dissolve in Certain Solvents?

Water is often called a universal solvent because many substances dissolve in it. However, no one substance can dissolve every solute. A general rule in chemistry is that “like dissolves like.” This rule means that a solvent will dissolve substances that have similar molecular structures. 

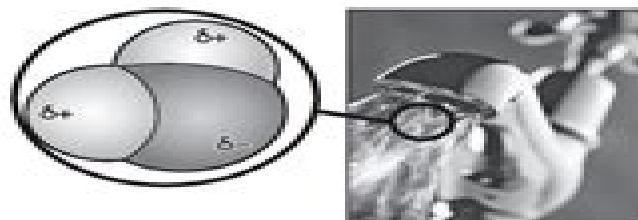
A solvent must be able to attract solute particles and pull them away from one another. Thus, solvent particles must attract solute particles more strongly than the solute particles attract one another.

POLAR MOLECULES

Recall that atoms bonded together do not all share electrons equally. For example, a water molecule is made up of two hydrogen atoms bonded to one oxygen atom. The oxygen atom attracts electrons more strongly than the hydrogen atoms do. Thus, the electrons spend more time near the oxygen atom. This gives each hydrogen atom a partial positive charge (δ^+) and the oxygen atom a partial negative charge (δ^-). 

Molecules that have partially charged positive and negative areas are **polar**. Because water molecules are polar, water can dissolve many other polar substances. Water can also dissolve ionic compounds such as table salt.

Water is a polar molecule because the oxygen atom strongly attracts electrons. Because the electrons spend more time near the oxygen atom, the oxygen atom has a partial negative charge.



READING TOOLBOX

Summarize Read each page of the section silently to yourself. With a partner, take turns summarizing each page.

READING CHECK

1. Identify Why is water called a universal solvent?

because
many substances dissolve
in it

READING CHECK

2. Explain Why does a water molecule have partial charges?

The oxygen atom attracts
electrons more strongly
than the hydrogen atoms
do. Thus, the electrons
spend more time near the
oxygen atom.

How Substances Dissolve Answer Key

BM King



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characteristics of water and water bodies water sciences water pollution and water wastewater treatment Part IV deals with soil science and emphasizes soil as a natural resource highlighting the many interactions between soil and other components of the ecosystem Part V is devoted to showing how decisions regarding handling solid and hazardous waste have or can have profound impact on the environment and the three media discussed in this text air water and soil Finally the epilogue looks at the state of the environment past present and future The emphasis in this brief unit is on mitigating present and future environmental concerns by incorporating technology into the remediation process not by blaming technology for the problem

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Table of Contents **How Substances Dissolve Answer Key**

1. Understanding the eBook **How Substances Dissolve Answer Key**
 - The Rise of Digital Reading **How Substances Dissolve Answer Key**
 - Advantages of eBooks Over Traditional Books
2. Identifying **How Substances Dissolve Answer Key**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **How Substances Dissolve Answer Key**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **How Substances Dissolve Answer Key**
 - Personalized Recommendations
 - **How Substances Dissolve Answer Key** User Reviews and Ratings
 - **How Substances Dissolve Answer Key** and Bestseller Lists

5. Accessing How Substances Dissolve Answer Key Free and Paid eBooks
 - How Substances Dissolve Answer Key Public Domain eBooks
 - How Substances Dissolve Answer Key eBook Subscription Services
 - How Substances Dissolve Answer Key Budget-Friendly Options
6. Navigating How Substances Dissolve Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - How Substances Dissolve Answer Key Compatibility with Devices
 - How Substances Dissolve Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of How Substances Dissolve Answer Key
 - Highlighting and Note-Taking How Substances Dissolve Answer Key
 - Interactive Elements How Substances Dissolve Answer Key
8. Staying Engaged with How Substances Dissolve Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers How Substances Dissolve Answer Key
9. Balancing eBooks and Physical Books How Substances Dissolve Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection How Substances Dissolve Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine How Substances Dissolve Answer Key
 - Setting Reading Goals How Substances Dissolve Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of How Substances Dissolve Answer Key
 - Fact-Checking eBook Content of How Substances Dissolve Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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