#### CK-12 Chemistry - Basic Answer Key

#### Chapter 7: Chemical Nomenclature

#### 7.1 Ionic Compounds

#### Lesson Review Questions

Owestions:

- What is the purpose of an empirical formula?
- Define binary and temary ionic compounds.
- What is the difference between a monatomic and a polyatomic ion?
- 4. For each of the following ionic compounds, what ions are present and in what ratio?
  - (a) MoBrs
  - (b) Li<sub>2</sub>CO<sub>3</sub>
  - (c) Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>
- Predict the compound that forms when gallium combines with oxygen. What would the name of this compound be?
- Give the formula for each of the following ionic compounds:
  - (a) ammonium nitrate

(d) variadium(III) oxide

(b) cobalt(II) sulfate

(e) barium oxide

(c) nickel(II) cyanide

- (f) calcium hypochlorite
- Name the following ionic compounds:
  - (a) MgBr<sub>2</sub>
  - (b) (d) KMinO<sub>4</sub>
  - (c) LivOOs
  - (d) (NH<sub>4</sub>)<sub>5</sub>S
  - (e) KHSO<sub>3</sub>
  - (f) CuCl (g) CuCl<sub>2</sub>
- White the correct formulas for the following ionic compounds:
  - a Barium chloride.

  - b. Chromium(III) oxide
  - Potassium sulfate
  - d. Zinc phosphate

Answer

 The empirical formula is the most straightforward way to describe the structure of a molecule, using a chemical formula that gives the lowest whole number ratio between the ions involved.

# **Mastering Chemistry Answer Key Chapter 7**

Robert C. Smoot, Smoot, Richard G. Smith, Jack Price

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