

JHS KAP CHEMISTRY – IONIC BONDING NOTES

CHEMICAL BONDING OVERVIEW

The focus of this unit is on the **IONIC BOND** and naming **IONIC** compounds which are compounds involving atoms bonded together by ionic bonds. We will briefly introduce metallic and covalent bonds as well.

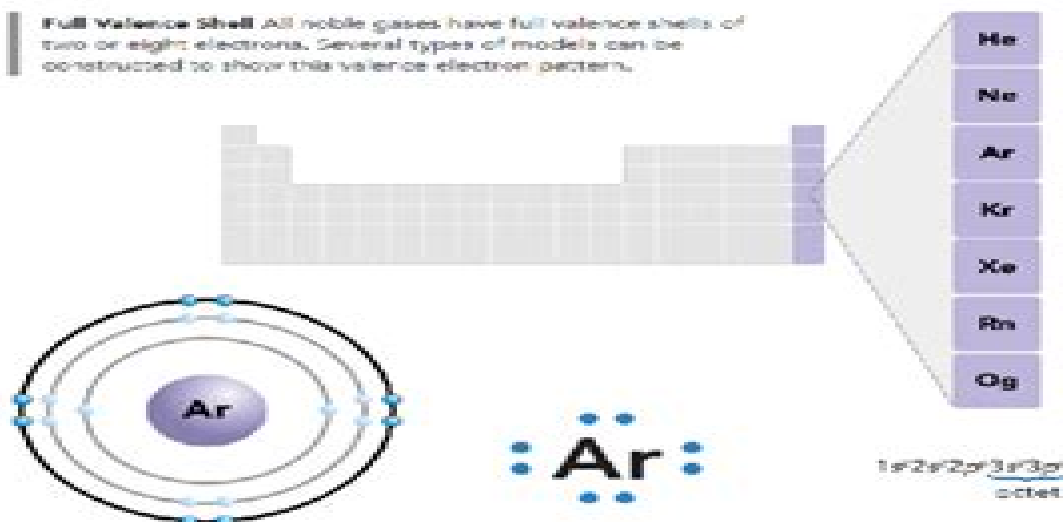
Chemical Bond

A chemical bond is a force of attraction that holds atoms or ions together.

The Octet Rule

- The **octet rule** states that atoms form chemical bonds so that each atom has **eight electrons** (an octet) in its **valence shell** (highest occupied energy level).
- Except for helium which has a total of two electrons, both in the valence shell, the **noble gases** each have **eight electrons** in their **valence shell**. This is an **energetically stable** valence shell electron configuration.
- Atoms of **other elements** achieve the **stable electron configuration** of noble gases by **forming chemical bonds** so that each atom ends up with an **octet** of electrons in their valence shell.
- A few atoms like **hydrogen**, **lithium**, and **beryllium** will become **stable with two electrons** in their valence shell like **helium**.

Full Valence Shell All noble gases have full valence shells of two or eight electrons. Several types of models can be constructed to show this valence electron pattern.



The Bohr model of an argon atom clearly shows that it has eight valence electrons. It also shows that inner electrons.

The electron dot structure shows only the eight valence electrons.

The electron configuration shows that two electrons in the 3s subshell and six in the 3p subshell make up the octet.

DEFINITIONS

If a substance is a **pure solid**, we use the symbol **(s)**. If it is a **pure liquid**, we use the symbol **(l)**. If it is a **pure gas**, we use the symbol **(g)**. When a substance **dissolves in water**, (is soluble in water), we use the symbol **(aq)** to show that it has formed an **aqueous solution**.

Examples:

Solids:	Na_2O	$\text{NaCl}_{(s)}$	$\text{CaCO}_{3(s)}$
Liquids:	$\text{H}_2\text{O}_{(l)}$	$\text{Br}_{2(l)}$	$\text{C}_2\text{H}_5\text{OH}_{(l)}$
Gases:	$\text{CO}_{2(g)}$	$\text{H}_{2(g)}$	$\text{NH}_{3(g)}$
Aqueous solutions:	$\text{LiF}_{(aq)}$	$\text{KNO}_{3(aq)}$	$\text{C}_6\text{H}_{12}\text{O}_{6(aq)}$

See the flow chart on the next page for more definitions.

Ionic Bonding Notes Guide

F Rizvi



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Table of Contents Ionic Bonding Notes Guide

1. Understanding the eBook Ionic Bonding Notes Guide
 - The Rise of Digital Reading Ionic Bonding Notes Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonding Notes Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonding Notes Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonding Notes Guide
 - Personalized Recommendations
 - Ionic Bonding Notes Guide User Reviews and Ratings
 - Ionic Bonding Notes Guide and Bestseller Lists
5. Accessing Ionic Bonding Notes Guide Free and Paid eBooks
 - Ionic Bonding Notes Guide Public Domain eBooks
 - Ionic Bonding Notes Guide eBook Subscription Services
 - Ionic Bonding Notes Guide Budget-Friendly Options

6. Navigating Ionic Bonding Notes Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Bonding Notes Guide Compatibility with Devices
 - Ionic Bonding Notes Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonding Notes Guide
 - Highlighting and Note-Taking Ionic Bonding Notes Guide
 - Interactive Elements Ionic Bonding Notes Guide
8. Staying Engaged with Ionic Bonding Notes Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonding Notes Guide
9. Balancing eBooks and Physical Books Ionic Bonding Notes Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonding Notes Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ionic Bonding Notes Guide
 - Setting Reading Goals Ionic Bonding Notes Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic Bonding Notes Guide
 - Fact-Checking eBook Content of Ionic Bonding Notes Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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