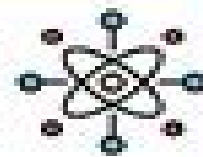


CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, creating oppositely charged ions that attract each other. Example: $NaCl$, MgO .
- Metallic Bond:** Formed by the attraction between metal cations and a sea of delocalized electrons. Example: Fe , Cu , Al .

Covalent, Ionic, and Metallic Bonding

Covalent bonding involves the sharing of electron pairs between atoms. This results in a stable molecule where the shared electrons are attracted to both nuclei. Covalent bonds are typically found in nonmetals and organic compounds.

Ionic bonding involves the complete transfer of one or more electrons from one atom to another. This results in the formation of positively charged cations and negatively charged anions, which are held together by strong electrostatic forces. Ionic bonds are typically found in salts and ionic compounds.

Metallic bonding involves the attraction between metal cations and a sea of delocalized electrons. The electrons are free to move throughout the metal lattice, which gives metals their characteristic properties, such as conductivity and malleability.

Science Real

Diagram of Ionic Bonding: Shows a red circle with a minus sign ($-$) and a yellow circle with a plus sign ($+$) connected by a dashed line, representing an ionic bond. Below it, a more complex structure shows multiple ions interacting.

Diagram of Covalent Bonding: Shows two atoms, each with a nucleus and electron shells. A red arrow points to the overlapping region of the outer shells, where electrons are shared. Below it, a more complex structure shows multiple atoms interacting.

Diagram of Metallic Bonding: Shows a lattice of metal cations (represented by blue circles with plus signs) surrounded by a sea of delocalized electrons (represented by a blue cloud). Below it, a more complex structure shows the arrangement of ions and electrons in a metal lattice.

Text on the right:

...forming a bond in a diatomic molecule. If each atom shares an electron with the other, both atoms have eight valence electrons.

...single, double, or triple covalent bonds. Triple and double bonds are stronger and shorter.

GUIDED READING ACTIVITY

Ionic Bonds Guided Reading And Study Answers

Wenbin Ji



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