

- and used in surgical instrument
4. Tin lowers the melting point of the alloy and is used for soldering purpose. D. Stainless steel
5. The alloy is hard, brittle, takes up polish and is used for making statues E. Solder

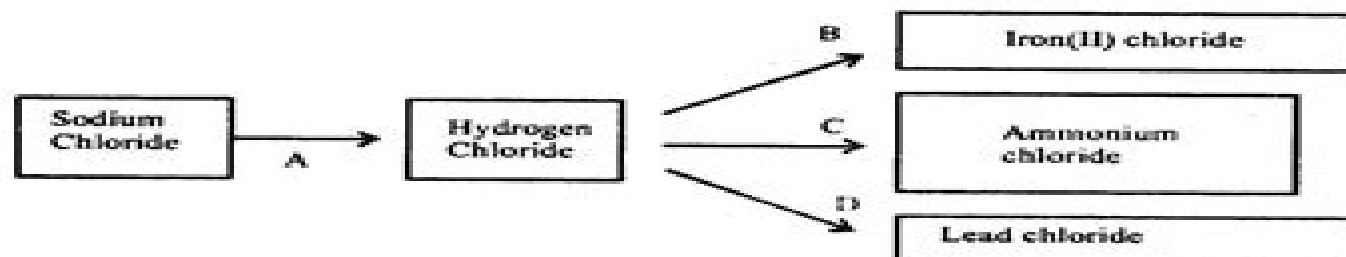
[5]

Question 4

- (a) Identify the anion present in the following compounds:
- Compound X on heating with copper turnings and concentrated sulphuric acid liberates a reddish brown gas.
 - When a solution of compound Y is treated with silver nitrate solution a white precipitate is obtained which is soluble in excess of ammonium hydroxide solution.
 - Compound Z which on reacting with dilute sulphuric acid liberates a gas which turns lime water milky, but the gas has no effect on acidified potassium dichromate solution.
 - Compound L on reacting with Barium chloride solution gives a white precipitate insoluble in dilute hydrochloric acid or dilute nitric acid. [4]
- (b) State one chemical test between each of the following pairs:
- Sodium carbonate and Sodium sulphite
 - Ferrous nitrate and Lead nitrate
 - Manganese dioxide and Copper (II) oxide [3]
- (c) Draw an electron dot diagram to show the structure of hydronium ion. State the type of bonding present in it. [3]

Question 5

- (a) (i) 67.2 litres of hydrogen combines with 44.8 litres of nitrogen to form ammonia under specific conditions as:
- $$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \longrightarrow 2\text{NH}_3(\text{g})$$
- Calculate the volume of ammonia produced. What is the other substance, if any, that remains in the resultant mixture? [2]
- (ii) The mass of 5.6 dm³ of a certain gas at STP is 12.0 g. Calculate the relative molecular mass of the gas. [2]
- (iii) Find the total percentage of Magnesium in magnesium nitrate crystals, Mg(NO₃)₂·6H₂O. [Mg = 24; N = 14; O = 16 and H = 1] [2]
- (b) Refer to the flow chart diagram below and give balanced equations with conditions, if any, for the following conversions A to D. [4]



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