

# Experiment 29 Prelaboratory Assignment

## Bleach Analysis

Date \_\_\_\_\_ Lab Sec. \_\_\_\_\_ Name \_\_\_\_\_ Desk No. \_\_\_\_\_

- Potassium iodate,  $\text{KIO}_3$ , and the bleach's hypochlorite ion,  $\text{ClO}^-$ , are oxidizing agents. What do they oxidize in this experiment?
  - Sodium thiosulfate,  $\text{Na}_2\text{S}_2\text{O}_3$ , is a reducing agent. What does it reduce in this experiment?
- Write the balanced net-ionic equation for the standardization of the sodium thiosulfate solution by combining equations 29.7 and 29.4.
- Experimental Procedure, Part A.1. Describe the procedure for the preparation of 100 mL of a 0.010 M  $\text{KIO}_3$  solution, starting with solid  $\text{KIO}_3$  (molar mass = 214.02 g/mol).
  - Experimental Procedure, Part B.1. Describe the procedure for the preparation of 250 mL of 0.10 M  $\text{Na}_2\text{S}_2\text{O}_3$  starting with solid  $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$  (molar mass = 248.2 g/mol).
- Experimental Procedure, Part B.3. Explain why a minimum, and not a maximum, mass of ~2 g of KI is critical for the standardization of the sodium thiosulfate solution.
- Write the balanced net-ionic equation for the analysis of the "available chlorine" in a sample using sodium thiosulfate as the reducing agent by combining equations 29.2, 29.3, and 29.4.



# General Chemistry Pre Lab Manual Answers

**WJ Hussar**



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