

Orbitals and Bonding Test #12 SCH4U Name: _____

 K T C A
15 10 8 6

Assessment of Learning (Product)

Multiple Choice: Circle the most correct answer. (15K)

1. What is the electron configuration of potassium?

- A) $1s^2 2s^2 2p^2 3s^2 3p^2 4s^2$ C) $1s^2 2s^2 3s^2 3p^6 3d^2$
B) $1s^2 2s^2 2p^{10} 3s^2 3p^3$ D) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$

2. Which electron configuration represents a violation of the Pauli exclusion principle?



3. Which one of the following sets of quantum numbers can correctly represent a 3p orbital?

- | | | | | |
|------------|------------|------------|------------|------------|
| A) $n = 3$ | B) $n = 1$ | C) $n = 3$ | D) $n = 3$ | E) $n = 3$ |
| $l = 1$ | $l = 3$ | $l = 2$ | $l = 1$ | $l = 0$ |
| $m_l = 2$ | $m_l = 3$ | $m_l = 1$ | $m_l = -1$ | $m_l = 1$ |

4. How many electrons in an atom can have $n = 4$, $l = 2$?

- A) 6 B) 14 C) 10 D) 2 E) None of these

General Chemistry II Chapter Covalent Bonding Orbitals

Practice Test

AW Rasmussen



General Chemistry II Chapter Covalent Bonding Orbitals Practice Test:

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