

Orbitals and Bonding Test #12 SCH4U Name: _____

 K T C A
 15 10 8 6

Assessment of Learning (Product)

Multiple Choice: Circle the most correct answer. (15K)

1. What is the electron configuration of potassium?
 A) $1s^2 2s^2 2p^2 3s^2 3p^2 4s^2$ C) $1s^2 2s^2 3s^2 3p^6 3d^2$
 B) $1s^2 2s^2 2p^{10} 3s^2 3p^3$ D) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$

2. Which electron configuration represents a violation of the Pauli exclusion principle?

A)

1s
↑

2s
↑↓

2p		

B)

1s
↑↑

2s
↑↓

2p		

C)

1s
↑↓

2s
↑↓

2p		
↑↓		

D)

1s
↑↓

2s
↑↓

2p		
↑		↑

E)

1s
↑

2s
↑

2p		
↑	↑	↑

3. Which one of the following sets of quantum numbers can correctly represent a 3p orbital?
 A) $n = 3, l = 1, m_l = 2$ B) $n = 1, l = 3, m_l = 3$ C) $n = 3, l = 2, m_l = 1$ D) $n = 3, l = 1, m_l = -1$ E) $n = 3, l = 0, m_l = 1$

4. How many electrons in an atom can have $n = 4, l = 2$?
 A) 6 B) 14 C) 10 D) 2 E) None of these

General Chemistry II Chapter Covalent Bonding Orbitals Practice Test

AW Rasmussen



General Chemistry II Chapter Covalent Bonding Orbitals Practice Test:

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