

Which of these are **true** and which are **false** about acid-base reactions?

- In general, an acid-base reaction will be favored when a weaker acid and a weaker base combine to form a stronger acid and a stronger base
Other way around: will be favored when a stronger acid and a stronger base combine to give a weaker acid and a weaker base.
T F
- Generally, strong acids have weak conjugate bases
True - and weak acids have strong conjugate bases
T F
- Strong acids tend to have high pK_a values
Strong acids have low (strongly negative) pK_a values - e.g. HCl has a pK_a of -8
T F
- The greater the acidity of a molecule, the weaker a base it is
Not quite correct. The stronger the acid, the weaker is its conjugate base.
T F
- The inductive effect of electronegative atoms near a carboxylic acid group increases the acidity of the compound.
T F
- Any factor which stabilizes negative charge generally tends to decrease the acidity of an acid
False - increases acidity, not decreases.
T F
- The pK_a of NH_4 is 10 and the pK_a of H_2O is 14, which means that H_2O is a stronger acid than H_2O
False - NH_4 is a stronger acid by a factor of 4 pK_a units (about 10^4)
T F
- The acidity of a molecule can be increased by resonance stabilization of its conjugate base.
T F
- sp -hybridized C-H bonds are more acidic than sp^3 -hybridized C-H bonds since the resulting negative charge is closer to the nucleus
T F
- Acid-base reactions where a stronger base is formed from a weaker base never occur.
They do occur and they can be important. E.g. NH_3 can deprotonate H_2O even though it is a weaker base by a factor of 10^4 compared to HO^- . Even a difference in 9-10 pK_a units can be significant for moving some reactions forward.
T F

General Chemistry Acid Base Practice Problems

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Table of Contents General Chemistry Acid Base Practice Problems

1. Understanding the eBook General Chemistry Acid Base Practice Problems
 - The Rise of Digital Reading General Chemistry Acid Base Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Acid Base Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Acid Base Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Acid Base Practice Problems

- Personalized Recommendations
- General Chemistry Acid Base Practice Problems User Reviews and Ratings
- General Chemistry Acid Base Practice Problems and Bestseller Lists
- 5. Accessing General Chemistry Acid Base Practice Problems Free and Paid eBooks
 - General Chemistry Acid Base Practice Problems Public Domain eBooks
 - General Chemistry Acid Base Practice Problems eBook Subscription Services
 - General Chemistry Acid Base Practice Problems Budget-Friendly Options
- 6. Navigating General Chemistry Acid Base Practice Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Acid Base Practice Problems Compatibility with Devices
 - General Chemistry Acid Base Practice Problems Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Acid Base Practice Problems
 - Highlighting and Note-Taking General Chemistry Acid Base Practice Problems
 - Interactive Elements General Chemistry Acid Base Practice Problems
- 8. Staying Engaged with General Chemistry Acid Base Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Acid Base Practice Problems
- 9. Balancing eBooks and Physical Books General Chemistry Acid Base Practice Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Acid Base Practice Problems
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Acid Base Practice Problems
 - Setting Reading Goals General Chemistry Acid Base Practice Problems
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Acid Base Practice Problems

- Fact-Checking eBook Content of General Chemistry Acid Base Practice Problems
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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