UNIT III - THERMOCHEMISTRY

Calculate the H-F bond energy given the following data.

 $H-H(g) + F-F(g) \rightarrow 2 H-F(g)$ $\Delta H_{rxn} = -544 \text{ kJ}$

A. 1130 kJ/mol B. 565 kJ/mol C.)283 J/mol D. 300 kJ/mol E. 134 kJ/mol

1.0 g of the rocket fuel hydrazine (N2H4) is burned in a bomb calorimeter. The temperature
of the bomb calorimeter changes from 24.63 °C to 28.16 °C. The heat capacity of the bomb
calorimeter is 837 J/K. Calculate the heat of combustion per mol of hydrazine in the bomb
calorimeter.

A. 95.0 kJ/mol B. 2.96 kJ/mol C. 3.55 x 10³ B. 1.13 x 10⁵ kJ/mol

Calculate the enthalpy change for the following reaction of nitric oxide NO with ozone O3.

 $NO(g) - O_3(g) \rightarrow NO_2(g) + O_2(g)$

Given the following data:

NO (g) + O (g) \rightarrow NO₂ (g) $\Delta H = -304.1 \text{ kJ}$ 2O₃ (g) \rightarrow 3O₂ (g) $\Delta H = -284.6 \text{ kJ}$ 2O (g) \rightarrow O₂ (g) $\Delta H = -495.0 \text{ kJ}$ A. -1083.7 kJ B. -693.8 kJ C. -198.8 kJ D -93.7 kJ E. -93.7 kJ

4. An endothermic reaction does no work on the surroundings. What values do you predict for the heat and energy change of this reaction?

A. q<0, ΔE>0

D. q>0, ΔE<0

E. q>0, ΔE=0

C. q>0, ΔE>0

 M&M candies contain 209.0. Calories per 1.5 oz serving. The fat content accounts for 38.3% of these Calories. What is the energy of the fat content in kJ?

A. 19.1 kJ B. 874 kJ S 1.91 x 10⁻² kJ D. 3.35 x 10⁵ kJ E. 335 kJ

Calculate the heat released when 2.00 g of butane from a cigarette lighter is burned.

2 C₄H₁₀ (l) + 13 O₂ (g) - 8 CO₂ (g) + 10 H₂O (l)

A. -90.7 kJ (B.)-98.3 kJ C. -49.2 kJ D. -108.4 kJ E. -196.6 kJ

7. When Karl Kaveman adds chilled grog to a granite mug, 10.9 kJ of energy is transferred from the mug to the grog. If the mug has a mass of 625 g and an initial temperature of 25 °C, what is the final temperature of the mug? The specific heat of granite is 0.79 J/g·K.

A. 2.9 °C B. 14 °C C. 22 °C D. 25 °C E. 47 °C

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