- Solution C1, an aqueous solution of potassium iodate (V) of concentration 4.06 g/dm³.
- Solution C2, an aqueous solution of sodium thiosulphate of unknown concentration.
- Solution C3, an acidified solution of potassium iodide.
- · Solution C4, starch indicator solution.

Section 1

You are required to:

- a) React solution C3 with C1 to liberate a certain quantity of iodine.
- b) Titrate the iodine liberated in the reaction between C3 and C1 against solution C2 to determine the molar concentration of C2.

Procedure

- (a) Fill the burette with solution C2.
- (b) Pipette 25.0 cm³ of solution C1 into a 250 ml conical flask.
- (c) Using a measuring cylinder, transfer 15.0cm³ of solution C3 into a conscal the conscal flask with solution C1.
- (d) Titrate the solution in the conical flask against solution C2 from the burette until the brown colour just changes to pale yellow.
- (e) Using a 10 ml measuring cylinder, add 5cm³ of solution C4 into the mixture in the conical flask and then continue titrating until the blue colour just disappears.
- (f) Record your titration results in table 1 below.
- (g) Repeat the procedure (b) to (f) above two more times and complete table 1 below.

Table 1 I II II

Final burette reading (cm³)

Initial burette reading (cm³)

Volume of solution C2 (cm³) added

(a) Determine the average volume of C2 used.

(I mark)

 Potassium iodate (V) solution reacts with acidified potassium iodide solution to liberate iodine as per the following ionic equation.

$$IO_3$$
 (aq) + 5Γ (aq) + $6H$ (aq) $\rightarrow 3I_2(aq) + 3H_2O$ (I)
(colourless) (brown)

(ii) Sodium thiosulphate reduces iodine to iodide ions as per the ionic equation below:

$$2S_2O_3^{2-}(aq) + I_2(aq) \rightarrow 2\Gamma(aq) + S_4O_6^{2-}(aq)$$
(Brown) (colourless)

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Vicenc Mendez, Sergei Fedotov, Werner Horsthemke

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