

THE MOLE: LAB ACTIVITY

Station 1:

A) Determine the mass of one drop of water by measuring the mass of 50 drops of water.

Mass of dish (or beaker)	=	_____	g
Mass of dish (or beaker) + 50 drops water	=	_____	g
Mass of 50 drops water	=	_____	g
Mass of 1 drop of water	=	_____	g

B) How many molecules are there in one drop of water?

Remember 1 mole of a substance is 6.02×10^{23} particles (atoms or molecules).
Therefore, 6.02×10^{23} molecules of water will weigh 18.02 g/mol

- Find # of moles of water for 1 drop $n = (\text{mass})/(\text{Molar Mass})$
- # of molecules in one drop of water $= n_{\text{water}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}}$
 $= ? \text{ molecules}$

Station 2

Calculate the number of copper atoms in a penny.

- Find # of moles of Cu $n = (\text{mass})/(\text{Molar Mass})$
- # of Cu atoms in a penny $= n_{\text{Cu}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$
 $= ? \text{ atoms}$

Station 3

A. Calculate the number of moles in one spoon of sugar. Assume that sugar is made up entirely of sucrose molecules ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$).

$$M_{\text{sucrose}} = 342.34 \text{ g/mol}$$

$$n_{\text{sugar}} = \frac{M_{\text{sucrose}}}{M_{\text{sucrose}}}$$

$$= \underline{\hspace{2cm}} \text{ mol}$$

B. Calculate the number of molecules of sugar in one spoon of sugar.

$$\begin{aligned} \# \text{ of molecules sucrose} &= n_{\text{sucrose}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= \underline{\hspace{2cm}} \text{ mol} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}} \\ &= ? \text{ molecules} \end{aligned}$$

Mole Lab Student Guide Answer Key

SJ Ball



Mole Lab Student Guide Answer Key:

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Table of Contents Mole Lab Student Guide Answer Key

1. Understanding the eBook Mole Lab Student Guide Answer Key
 - The Rise of Digital Reading Mole Lab Student Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Mole Lab Student Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mole Lab Student Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Lab Student Guide Answer Key
 - Personalized Recommendations
 - Mole Lab Student Guide Answer Key User Reviews and Ratings
 - Mole Lab Student Guide Answer Key and Bestseller Lists

5. Accessing Mole Lab Student Guide Answer Key Free and Paid eBooks
 - Mole Lab Student Guide Answer Key Public Domain eBooks
 - Mole Lab Student Guide Answer Key eBook Subscription Services
 - Mole Lab Student Guide Answer Key Budget-Friendly Options
6. Navigating Mole Lab Student Guide Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Mole Lab Student Guide Answer Key Compatibility with Devices
 - Mole Lab Student Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mole Lab Student Guide Answer Key
 - Highlighting and Note-Taking Mole Lab Student Guide Answer Key
 - Interactive Elements Mole Lab Student Guide Answer Key
8. Staying Engaged with Mole Lab Student Guide Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole Lab Student Guide Answer Key
9. Balancing eBooks and Physical Books Mole Lab Student Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole Lab Student Guide Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mole Lab Student Guide Answer Key
 - Setting Reading Goals Mole Lab Student Guide Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mole Lab Student Guide Answer Key
 - Fact-Checking eBook Content of Mole Lab Student Guide Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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