

Molar Mass and Mole Calculations Worksheet

1. Calculate the molar mass for each of the following compounds:

- | | |
|--|---|
| a. $\text{NaCl} = 58.44 \text{ g/mol}$ | e. $(\text{NH}_4)_3\text{PO}_4 = 149.12 \text{ g/mol}$ |
| b. $\text{Fe}_2\text{O}_3 = 159.7 \text{ g/mol}$ | f. $\text{Ca}_3(\text{PO}_4)_2 = 310.18 \text{ g/mol}$ |
| c. $\text{MgCO}_3 = 84.32 \text{ g/mol}$ | g. $\text{C}_6\text{H}_{12}\text{O}_6 = 180.18 \text{ g/mol}$ |
| d. $\text{Ba}(\text{OH})_2 = 171.35 \text{ g/mol}$ | h. $\text{C}_5\text{H}_5\text{N} = 79.11 \text{ g/mol}$ |

2. Calculate the number of moles in each of the following masses:

- a. 45.0 g of acetic acid, $\text{CH}_3\text{COOH} = 0.75 \text{ mol}$
- b. 7.04 g of lead (II) nitrate = 0.02 mol

3. Calculate the mass in grams of each of the following amounts:

- a. 3.00 mol of selenium oxybromide, $\text{SeOBr}_2 = 764.31 \text{ g}$
 $M = 78.97 + 16 + 2(79.9) = 254.77 \text{ g/mol} \quad 254.77 \times 3 =$
- b. 488 mol of calcium carbonate = 48,843.92 g
 $M = 40.08 + 12.01 + 3(16) = 100.09 \text{ g/mol} \quad 100.09 \times 488 =$

4. Calculate the number of molecules or formula units in each of the following amounts:

- a. 4.27 mol of tungsten(VI) oxide, $\text{WO}_3 = 2.57 \times 10^{24}$
 $4.27 \times 6.02 \times 10^{23} =$
- b. 0.989 mol of nitrobenzene $\text{C}_6\text{H}_5\text{NO}_2 = 5.95 \times 10^{23}$
 $0.989 \times 6.02 \times 10^{23} =$

5. Calculate the mass of each of the following quantities:

- a. 8.39×10^{23} molecules of fluorine = 52.96 g
 $8.39 \times 10^{23} / 6.02 \times 10^{23} = 1.39 \text{ mol} \quad 1.39 \times 38 =$
 $M = 2(19) = 38 \text{ g/mol}$
- b. 6.82×10^{24} formula units of beryllium sulfate = 1,190.33 g
 $M = 9.01 + 32.06 + 4(16) = 105.07 \quad 105.07 \times 11.33 =$
 $6.82 \times 10^{24} / 6.02 \times 10^{23} = 11.33$

Molar Mass And Chemical Quantities Answer Key

Ian Pickup



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