

CHEMISTRY: LIMITING REAGENT Practice Problems

Please show your work on a separate piece of paper!

Assume standard temperature and pressure, STP conditions for all reactions.

1. At high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide:



In one experiment, 7.62 g of Fe are allowed to react with 8.67 g of S.

- What is the limiting reagent, and what is the reactant in excess?
- Calculate the mass of FeS formed.

2. Acrylonitrile, $\text{C}_3\text{H}_3\text{N}$, is the starting material for the production of a kind of synthetic fiber (acrylics) and can be made from propylene, C_3H_6 , by reaction with nitric oxide, NO , as follows:



What mass of $\text{C}_3\text{H}_3\text{N}$ can be made when 21.6 g of C_3H_6 react with 21.6 g of nitric oxide?

3. Calculate the percent yield for the reaction: $\text{P}_{4(s)} + 6 \text{Cl}_{2(g)} \rightarrow 4 \text{PCl}_{3(l)}$
if 75.0 g of phosphorus reacts with excess chlorine gas to produce 111.0 g of phosphorus trichloride.

4. Formic acid, HCHO_2 , burns in oxygen to form carbon dioxide and water as follows:



If a 3.15-g sample of formic acid was burned in 2.0 L of oxygen, what volume of carbon dioxide would be produced?

5. Zinc metal reacts with hydrochloric acid to produce zinc chloride and hydrogen gas.
- Balance the following reaction: $\text{Zn}_{(s)} + \text{HCl}_{(aq)} \rightarrow \text{ZnCl}_{2(aq)} + \text{H}_{2(g)}$
 - A 3.50-g sample of zinc metal is allowed to react with 2.50 g of hydrochloric acid.

Complete the following table:

reactants/products	Zn (grams)	HCl (grams)	ZnCl ₂ (grams)	H ₂ (L)
Before reaction				
After reaction	1.26 g			

Limiting Reactant Problems Answer Key

James E. Brady, Fred Senese



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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

ChemDiscovery Teacher Edition Olga I. Agapova, 2002 **Mapping College Chemistry** Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this competitive environment entails not just studying harder or longer it means re organizing textbook content so that it is meaningful to the student This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career

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Investigating Chemistry Matthew Johl,2008-12-22 In its new second edition Investigating Chemistry A Forensic Science Perspective remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for

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Basics for Chemistry David A. Ucko, 2013-09-24 Basics of Chemistry provides the tools needed in the study of General Chemistry such as problem solving skills calculation methods and the language and basic concepts of chemistry The book is designed to meet the specific needs of underprepared students Concepts are presented only as they are needed and developed from the simple to the complex The text is divided into 18 chapters each covering some particular aspect of chemistry such as matter energy and measurement the properties of atoms description of chemical bonding study of chemical change and nuclear and organic chemistry Undergraduate students will find the book as a very valuable academic material

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