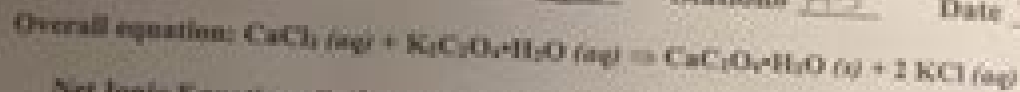


Name Mikayla D'AnnaSection B54Station A5Date 10/14/19

1. Data & Calculations:

CaCl₂ Concentration of solution <u>0.75 mol/L</u> Volume of solution <u>15 mL</u> Moles of CaCl₂ used <u>5.92×10^{-4} mol</u>	K₂C₂O₄·H₂O Mass assigned <u>0.140 g</u> Mass used <u>0.142 g</u> Moles of K₂C₂O₄·H₂O used <u>0.0008 mol</u> (MW = 184.14 g/mol)
Crucible and filter mass Dry (crucible + filter) <u>22.115 g</u> With Precipitate <u>22.221 g</u>	CaC₂O₄·H₂O (Precipitate) Mass of CaC₂O₄·H₂O recovered <u>0.106 g</u> Moles of CaC₂O₄·H₂O recovered <u>0.0007 mol</u> (MW = 146.12 g/mol)

Be sure to include the necessary units for the values reported in the above table. Attach sample calculations of each type necessary to obtain the results in the above table. These should include the equations used, substituted values, appropriate units, and results.

2. Excess and Limiting Tests:

2a. Did a precipitate form (1) when K₂C₂O₄ solution was added to the filtrate? yes (2) when CaCl₂ solution was added? No

2b. Based on the results shown in 2a, what was the limiting reactant? K₂C₂O₄ What was the excess reactant? CaCl₂

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Christian G. Meyer



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