

p. 548 (1-5, 10-18, 24-29, 29)
p. 561 (1, 4, 7, 8)

① Equilateral Δ



$$\begin{aligned} m\angle 1: \frac{360}{3} &= 120^\circ \\ m\angle 2: \frac{1}{2}(120) &= 60^\circ \\ m\angle 3: \frac{1}{2}(60) &= 30^\circ \end{aligned}$$

③ Regular Hexagon:

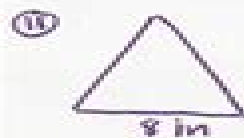


$$\begin{aligned} m\angle 1 &= \frac{360}{6} = 60^\circ \\ m\angle 2 &= \frac{1}{2}(60) = 30^\circ \\ m\angle 3 &= 60^\circ \end{aligned}$$



$$\begin{aligned} 18 \text{ ft } A &= \frac{1}{2} P a \\ A &= \frac{1}{2}(108)(9\sqrt{3}) \\ &= \boxed{486\sqrt{3} \text{ ft}^2} \end{aligned}$$

$$\begin{aligned} P: 18 \cdot 6 &= 108 \\ a: 9\sqrt{3} \end{aligned}$$



$$A = \frac{8^2\sqrt{3}}{4} = \boxed{16\sqrt{3} \text{ in}^2}$$



$$\begin{aligned} A &= \frac{1}{2} P a \\ &= \frac{1}{2}(36)(3\sqrt{3}) \\ &= \boxed{54\sqrt{3} \text{ m}^2} \end{aligned}$$

$$\begin{aligned} P: 6 \cdot 6 &= 36 \text{ in} \\ a: 3\sqrt{3} \end{aligned}$$



$$\begin{aligned} \text{So, } \frac{S}{2} &= 6.35\sqrt{3} \\ 8 &= 12.7\sqrt{3} \end{aligned}$$

$$A = \frac{S^2\sqrt{3}}{4} = \frac{(12.7\sqrt{3})^2\sqrt{3}}{4} = \frac{483.87\sqrt{3}}{4} = 209.52 \approx \boxed{210 \text{ in}^2}$$

② Square:



$$\begin{aligned} m\angle 4: 90^\circ \\ m\angle 5: 45^\circ \\ m\angle 6: 45^\circ \end{aligned}$$



$$\begin{aligned} A &= L \cdot W \\ (6\sqrt{2})(6\sqrt{2}) &= 36 \cdot 2 = \boxed{72 \text{ cm}^2} \end{aligned}$$



$$\begin{aligned} P: 16 \cdot 6 &= 96 \\ a: 8\sqrt{3} \end{aligned} \quad A = \frac{1}{2} P a$$

$$\frac{1}{2}(96)(8\sqrt{3}) = \boxed{384\sqrt{3} \text{ in}^2}$$

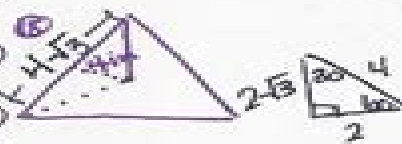


$$\begin{aligned} P: 6\sqrt{3} \cdot 6 &= 36\sqrt{3} \\ a: 9 \end{aligned} \quad A = \frac{1}{2} P a$$

$$\frac{1}{2}(36\sqrt{3})(9) = \boxed{162\sqrt{3} \text{ m}^2}$$



$$\begin{aligned} 10\sqrt{3} \cdot 5 \\ 5\sqrt{3} \end{aligned} \quad A = \frac{S^2\sqrt{3}}{4} = \frac{(10\sqrt{3})^2\sqrt{3}}{4} = \frac{300\sqrt{3}}{4} = \boxed{75\sqrt{3} \text{ m}^2}$$



$$A = \frac{S^2\sqrt{3}}{4} = \frac{(4\sqrt{3})^2\sqrt{3}}{4} = \frac{48\sqrt{3}}{4} = \boxed{12\sqrt{3} \text{ in}^2}$$

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web when you are trying to draw a lewis structure there are guidelines to do it number of steps can be changed according to the complexity of the molecule or ion because p2o5 molecule is a bit complex molecule if you are a beginner to lewis structure drawing you have to be careful when it starts to draw

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web dec 13 2017 step 1 draw the lewis structure there are three central atoms p p o arrange them symmetrically in the order p o p attach the remaining o atoms as terminal atoms and arrange electrons to give each atom an octet the p atoms will get expanded octets step 2 use vsepr theory to determine the electron geometries

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web jun 27 2022 a lewis electron dot diagram or electron dot diagram or a lewis diagram or a lewis structure is a representation of the valence electrons of an atom that uses dots around the symbol of the element the number of dots equals the number of valence electrons in the atom

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web a lewis structure is a way to show the shape of a molecule dots show where electrons are around the atoms and lines or pairs of dots show where covalent bonds connect the atoms by drawing a lewis dot structure you can find the lone electron pairs in molecules which helps you figure out how chemical bonds form

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web aug 4 2019 step 1 find the total number of valence electrons iodine has 7 valence electrons chlorine has 7 valence electrons total valence electrons 1 iodine 7 3 chlorine 3 x 7 total valence electrons 7 21 total valence electrons 28 step 2 find the number of electrons needed to make the atoms happy iodine needs 8 valence electrons

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web a lewis diagram shows how the valence electrons are distributed around the atoms in a molecule shared pairs of electrons are drawn as lines between atoms while lone pairs of electrons are drawn as dots next to atoms

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web oct 8 2023 introduction lewis structure is basically a graphic representation of the electron distribution around an atom the major reason why learning lewis dot structure is important is that it helps in predicting the number and type of bonds which can be formed around an atom it also helps in predicting the geometry of the molecule