

Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)



- a. If 15.0g of N_2O_4 was produced, how many moles of O_2 were required?

$$\frac{15.0\text{g N}_2\text{O}_4}{92.0\text{g N}_2\text{O}_4} \times \frac{1\text{ mol N}_2\text{O}_4}{1\text{ mol N}_2\text{O}_4} \times \frac{2\text{ mol O}_2}{1\text{ mol N}_2\text{O}_4} = 0.326\text{ mol O}_2$$

- b. If 4.0×10^{-3} moles of oxygen reacted, how many grams of N_2 were needed?

$$\frac{4.0 \times 10^{-3}\text{ mol O}_2}{2\text{ mol O}_2} \times \frac{1\text{ mol N}_2}{1\text{ mol O}_2} \times \frac{28.0\text{g N}_2}{1\text{ mol N}_2} = 5.6 \times 10^{-2}\text{ g N}_2$$



- a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

$$\frac{2.04\text{ mol K}_3\text{PO}_4}{1\text{ mol K}_3\text{PO}_4} \times \frac{3\text{ mol KNO}_3}{1\text{ mol K}_3\text{PO}_4} \times \frac{101.1\text{ g KNO}_3}{1\text{ mol KNO}_3} = 619\text{g KNO}_3$$

- b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

$$\frac{5.80\text{g AlPO}_4}{122.0\text{g AlPO}_4} \times \frac{1\text{ mol AlPO}_4}{1\text{ mol AlPO}_4} \times \frac{1\text{ mol Al}(\text{NO}_3)_3}{1\text{ mol AlPO}_4} = 4.75 \times 10^{-2}\text{ mol Al}(\text{NO}_3)_3$$



- a. If you have 5.50 mol of CaC_2 , how much C_2H_2 do you get?

$$\frac{5.50\text{ mol CaC}_2}{1\text{ mol CaC}_2} \times \frac{1\text{ mol C}_2\text{H}_2}{1\text{ mol CaC}_2} \times \frac{26.0\text{g C}_2\text{H}_2}{1\text{ mol C}_2\text{H}_2} = 143\text{g C}_2\text{H}_2$$

- b. How many moles of water are needed when 65.0g of CaC_2 have reacted?

$$\frac{65.0\text{g CaC}_2}{64.1\text{g CaC}_2} \times \frac{1\text{ mol CaC}_2}{1\text{ mol CaC}_2} \times \frac{2\text{ mol H}_2\text{O}}{1\text{ mol CaC}_2} = 2.03\text{ mol H}_2\text{O}$$

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose ($\text{C}_6\text{H}_{12}\text{O}_6$). Write and balance the chemical equation. How many moles of CO_2 are required to make 120.0g of glucose?



$$\frac{120.0\text{g C}_6\text{H}_{12}\text{O}_6}{180.0\text{g C}_6\text{H}_{12}\text{O}_6} \times \frac{1\text{ mol C}_6\text{H}_{12}\text{O}_6}{1\text{ mol C}_6\text{H}_{12}\text{O}_6} \times \frac{6\text{ mol CO}_2}{1\text{ mol C}_6\text{H}_{12}\text{O}_6} = 4.000\text{ mol CO}_2$$

Mole Lab Answer Key

Victor M. Corman

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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hasn't a clue what to do with his semi stardom as a celebrity chef his parents have become swingers with whom is too shocking to go into now his epic novel is still unpublished his ex flame Pandora is running for political office and his younger sister has rebelled in the most distressingly common ways But there's one upside Adrian's son has inherited his mother's unblemished skin Is it any wonder that at 34 Adrian is still punishingly self aware and willfully deluded about what he's endured and what he's yet to achieve Struggling somewhere between breakthrough and breakdown he's telling his diary everything The result Adrian's fifth Book of Revelation and it's quite possibly a classic Daily Mirror

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