

Molar Mass and Mole Calculations Worksheet

1. Calculate the molar mass for each of the following compounds:

- | | |
|---|--|
| a. NaCl = 58.44 g/mol | e. $(\text{NH}_4)_3\text{PO}_4$ = 149.12 g/mol |
| b. Fe_2O_3 = 159.7 g/mol | f. $\text{Ca}_3(\text{PO}_4)_2$ = 310.18 g/mol |
| c. MgCO_3 = 84.32 g/mol | g. $\text{C}_6\text{H}_{12}\text{O}_6$ = 180.18 g/mol |
| d. $\text{Ba}(\text{OH})_2$ = 171.35 g/mol | h. $\text{C}_5\text{H}_5\text{N}$ = 79.11 g/mol |

2. Calculate the number of moles in each of the following masses:

- a. 45.0 g of acetic acid, CH_3COOH = **0.75 mol**
- b. 7.04 g of lead (II) nitrate = **0.02 mol**

3. Calculate the mass in grams of each of the following amounts:

- a. 3.00 mol of selenium oxybromide, SeOBr_2 = **764.31 g**
 $M = 78.97 + 16 + 2(79.9) = 254.77 \text{ g/mol}$ $254.77 \times 3 =$
- b. 488 mol of calcium carbonate = **48,843.92 g**
 $M = 40.08 + 12.01 + 3(16) = 100.09 \text{ g/mol}$ $100.09 \times 488 =$

4. Calculate the number of molecules or formula units in each of the following amounts:

- a. 4.27 mol of tungsten(VI) oxide, WO_3 = **2.57×10^{24}**
 $4.27 \times 6.02 \times 10^{23} =$
- b. 0.989 mol of nitrobenzene $\text{C}_6\text{H}_5\text{NO}_2$ = **5.95×10^{23}**
 $0.989 \times 6.02 \times 10^{23} =$

5. Calculate the mass of each of the following quantities:

- a. 8.39×10^{23} molecules of fluorine = **52.96 g**
 $8.39 \times 10^{23} / 6.02 \times 10^{23} = 1.39 \text{ mol}$ $1.39 \times 38 =$
 $M = 2(19) = 38 \text{ g/mol}$
- b. 6.82×10^{24} formula units of beryllium sulfate = **1,190.33 g**
 $M = 9.01 + 32.06 + 4(16) = 105.07$ $105.07 \times 11.33 =$
 $6.82 \times 10^{24} / 6.02 \times 10^{23} = 11.33$

Mole And Mass Answer Key

David Baud

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

Mole And Mass Answer Key:

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