#### Chemistry Worksheet: Limiting Reactant Worksheet #1

- 1. Consider the following reaction: 2 Al + 6 HBr → 2 AlBr<sub>3</sub> + 3 H<sub>2</sub>
  - a. When 3.22 moles of Al reacts with 4.96 moles of HBr, how many moles of H2 are formed?
  - b. What is the limiting reactant?
  - c. For the reactant in excess, how many moles are left over at the end of the
  - a) 4.83 mol Hz and 2.48 mol Hz
  - 6) HBr is the limiting readont
- 2. Consider the following reaction: 3 Si + 2 N<sub>2</sub> → Si<sub>3</sub>N<sub>4</sub>
  - a. When 21.44 moles of Si reacts with 17.62 moles of N<sub>2</sub>, how many moles of Si<sub>3</sub>N<sub>4</sub> are formed?
  - b. What is the limiting reactant?
  - c. For the reactant in excess, how many moles are left over at the end of the reaction?
  - Willicon is the limiting readant
  - c) 3, 34 mol Ne are left
- 3. Consider the following reaction: 2 CuCl<sub>2</sub> + 4 KI → 2 CuI + 4 KCl + I<sub>2</sub>
  - a. When 0.56 moles of CuCl<sub>2</sub> reacts with 0.64 moles of KI, how many moles of I<sub>2</sub> are formed?
  - b. What is the limiting reactant?
  - c. For the reactant in excess, how many moles are left over at the end of the reaction?

WO. 28 and O. 16 mol JZ b) KI is the limiting readant

UD. 24 mol CuClz are left

- Consider the following reaction: 4 FeS<sub>2</sub> + 11 O<sub>2</sub> → 2 Fe<sub>2</sub>O<sub>3</sub> + 8 SO<sub>2</sub>
  - a. When 26.62 moles of FeS<sub>2</sub> reacts with 5.44 moles of O<sub>2</sub>, how many moles of SO<sub>2</sub> are formed?
  - b. What is the limiting reactant?
  - c. For the reactant in excess, how many moles are left over at the end of the reaction?
     a) 53.24 and 3.95 mel 50z

502 1.27 mol Fe 52 are left

44

# **Limiting Reagent 1 Answer Key**

**ML Yell** 

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## Enjoying the Melody of Phrase: An Mental Symphony within Limiting Reagent 1 Answer Key

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