

Chapter 4 Atomic Structure

Section 4.2 The Structure of an Atom**(pages 108–112)**

This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.

Reading Strategy (page 108)

Monitoring Your Understanding Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Know About Atoms	What I Would Like to Learn	What I Have Learned
Most students will know that atoms are the “building blocks” of matter, and some may know that atoms contain subatomic particles.	Based on the title of the section, students may say that they want to learn more about the structure of atoms.	

Properties of Subatomic Particles (pages 108–109)

- What are three subatomic particles?
a. Protons b. Electrons c. Neutrons
- Circle the letter that identifies a subatomic particle with a positive charge.
a. nucleus **(b)** proton
c. neutron d. electron
- Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? A charged object did not deflect the paths of the particles.

Comparing Subatomic Particles (pages 109–110)

- Circle the letters of properties that vary among subatomic particles.
a. color **(b)** mass
(c) charge **(d)** location in the atom
- Circle the letter of the expression that accurately compares the masses of neutrons and protons.
(a) mass of 1 neutron = mass of 1 proton
b. mass of 2000 neutrons = mass of 1 proton
c. mass of 1 electron = mass of 1 proton
d. mass of 1 neutron = mass of 1 electron

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Andy Hewitt



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Table of Contents Hewitt Chapter 4 Worksheet Answers

1. Understanding the eBook Hewitt Chapter 4 Worksheet Answers
 - The Rise of Digital Reading Hewitt Chapter 4 Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Hewitt Chapter 4 Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hewitt Chapter 4 Worksheet Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Hewitt Chapter 4 Worksheet Answers
 - Personalized Recommendations
 - Hewitt Chapter 4 Worksheet Answers User Reviews and Ratings
 - Hewitt Chapter 4 Worksheet Answers and Bestseller Lists
5. Accessing Hewitt Chapter 4 Worksheet Answers Free and Paid eBooks
 - Hewitt Chapter 4 Worksheet Answers Public Domain eBooks
 - Hewitt Chapter 4 Worksheet Answers eBook Subscription Services
 - Hewitt Chapter 4 Worksheet Answers Budget-Friendly Options
6. Navigating Hewitt Chapter 4 Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Hewitt Chapter 4 Worksheet Answers Compatibility with Devices
 - Hewitt Chapter 4 Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hewitt Chapter 4 Worksheet Answers
 - Highlighting and Note-Taking Hewitt Chapter 4 Worksheet Answers
 - Interactive Elements Hewitt Chapter 4 Worksheet Answers
8. Staying Engaged with Hewitt Chapter 4 Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hewitt Chapter 4 Worksheet Answers
9. Balancing eBooks and Physical Books Hewitt Chapter 4 Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hewitt Chapter 4 Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Hewitt Chapter 4 Worksheet Answers
 - Setting Reading Goals Hewitt Chapter 4 Worksheet Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Hewitt Chapter 4 Worksheet Answers
 - Fact-Checking eBook Content of Hewitt Chapter 4 Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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