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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N_2 what is the limiting reagent if 8.52 g are actually formed what is the

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produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have 0.216232 grams/mole $\times$ 21.3 grams = 0.66 moles $\times$ 32 g/mol

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H₂ are left over 13

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield / actual yield $\times$ 100 based on this definition we would expect a percent yield

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