Part A

Calculate the amount of heat (in kilojoules) required to vaporize 2.58 kg of water at its boiling point.

Express the heat in kilojoules to three significant figures.

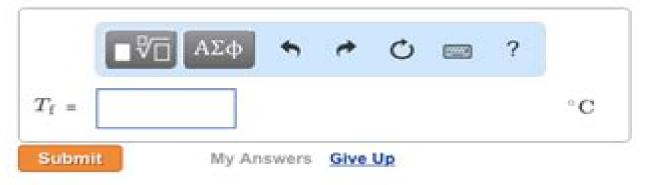
Correct

At its boiling point, the enthalpy of vaporization ($\Delta H_{\rm vap}$) of water is 40.7 kJ/mol. Once the given mass, 2.58×10^3 g, is converted to the number of moles using the molar mass of water (18.02 g/mol), multiplying the number of moles by $\Delta H_{\rm vap}$ yields the amount of heat required to evaporate that amount of water.

Part B

Suppose that 0.48 g of water at 25° C condenses on the surface of a 55-g block of aluminum that is initially at 25° C. If the heat released during condensation goes only toward heating the metal, what is the final temperature (in degrees Celsius) of the metal block? (The specific heat capacity of aluminum, $C_{s,A1}$, is 0.903 $J/(g \cdot {}^{\circ}C)$.)

Express the temperature in degrees Celsius to two significant figures.



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Jiyuan Zhang

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Table of Contents Mastering Chemistry Answers Chapter 12

- 1. Understanding the eBook Mastering Chemistry Answers Chapter 12
 - The Rise of Digital Reading Mastering Chemistry Answers Chapter 12
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Mastering Chemistry Answers Chapter 12
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Answers Chapter 12
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mastering Chemistry Answers Chapter 12
 - Personalized Recommendations
 - Mastering Chemistry Answers Chapter 12 User Reviews and Ratings
 - Mastering Chemistry Answers Chapter 12 and Bestseller Lists
- 5. Accessing Mastering Chemistry Answers Chapter 12 Free and Paid eBooks
 - Mastering Chemistry Answers Chapter 12 Public Domain eBooks
 - Mastering Chemistry Answers Chapter 12 eBook Subscription Services
 - Mastering Chemistry Answers Chapter 12 Budget-Friendly Options
- 6. Navigating Mastering Chemistry Answers Chapter 12 eBook Formats
 - o ePub, PDF, MOBI, and More
 - Mastering Chemistry Answers Chapter 12 Compatibility with Devices
 - Mastering Chemistry Answers Chapter 12 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Chemistry Answers Chapter 12
 - Highlighting and Note-Taking Mastering Chemistry Answers Chapter 12
 - Interactive Elements Mastering Chemistry Answers Chapter 12
- 8. Staying Engaged with Mastering Chemistry Answers Chapter 12

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Mastering Chemistry Answers Chapter 12
- 9. Balancing eBooks and Physical Books Mastering Chemistry Answers Chapter 12
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Answers Chapter 12
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mastering Chemistry Answers Chapter 12
 - Setting Reading Goals Mastering Chemistry Answers Chapter 12
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mastering Chemistry Answers Chapter 12
 - Fact-Checking eBook Content of Mastering Chemistry Answers Chapter 12
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - o Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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