

# Key

## Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PV = nRT$ ", and the universal gas constant  $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$  to solve the following problems:

If pressure is needed in kPa then convert by multiplying by  $101.3 \text{ kPa} / 1 \text{ atm}$  to get  
 $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of  $87^\circ\text{C}$ , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

$$\text{or } P = 166.29 \text{ kPa}$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of  $56^\circ\text{C}$ , what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 327 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

- 5) If I have 17 moles of gas at a temperature of  $67^\circ\text{C}$ , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 390 \text{ K}}{88.89 \text{ L}}$$

$$P = 5.34 \text{ atm}$$

$$\text{or } P = 540.61 \text{ kPa}$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

# Ideal Gas Law Problems Worksheet Answers

**M Lipman**



## **Ideal Gas Law Problems Worksheet Answers:**

## **Ideal Gas Law Problems Worksheet Answers** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the power of words has are more evident than ever. They have the capability to inspire, provoke, and ignite change. Such could be the essence of the book **Ideal Gas Law Problems Worksheet Answers**, a literary masterpiece that delves deep to the significance of words and their effect on our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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### **Table of Contents Ideal Gas Law Problems Worksheet Answers**

1. Understanding the eBook Ideal Gas Law Problems Worksheet Answers
  - The Rise of Digital Reading Ideal Gas Law Problems Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ideal Gas Law Problems Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ideal Gas Law Problems Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ideal Gas Law Problems Worksheet Answers
  - Personalized Recommendations
  - Ideal Gas Law Problems Worksheet Answers User Reviews and Ratings
  - Ideal Gas Law Problems Worksheet Answers and Bestseller Lists

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  - Ideal Gas Law Problems Worksheet Answers eBook Subscription Services
  - Ideal Gas Law Problems Worksheet Answers Budget-Friendly Options
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  - Ideal Gas Law Problems Worksheet Answers Compatibility with Devices
  - Ideal Gas Law Problems Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ideal Gas Law Problems Worksheet Answers
  - Highlighting and Note-Taking Ideal Gas Law Problems Worksheet Answers
  - Interactive Elements Ideal Gas Law Problems Worksheet Answers
8. Staying Engaged with Ideal Gas Law Problems Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ideal Gas Law Problems Worksheet Answers
9. Balancing eBooks and Physical Books Ideal Gas Law Problems Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ideal Gas Law Problems Worksheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ideal Gas Law Problems Worksheet Answers
  - Setting Reading Goals Ideal Gas Law Problems Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ideal Gas Law Problems Worksheet Answers
  - Fact-Checking eBook Content of Ideal Gas Law Problems Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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