

### Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility:** the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.  
**saturated solution:** a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution:** a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscibles:** describes liquids that dissolve in each other. **immiscible:** describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47.  $5.55 \times 10^2$  g  $\text{AgNO}_3$
48. Solubility increases with pressure.
49. a.  $1.6 \times 10^{-2}$  g/L  
b.  $4.7 \times 10^{-2}$  g/L
50. Dilute and concentrated are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.  
a. 1.3M KCl  
b.  $3.3 \times 10^{-1}$  M  $\text{MgCl}_2$
52.  $2.00 \times 10^3$  mL
53. a.  $5.0 \times 10^{-1}$  mol NaCl, 29 g NaCl  
b. 1.0 mol  $\text{KNO}_3$ ,  $1.0 \times 10^2$  g  $\text{KNO}_3$   
c.  $2.5 \times 10^3$  mol  $\text{CaCl}_2$ , 2.8 g  $\text{CaCl}_2$
54. a.  $2.3 \times 10^3$  g NaCl  
b. 2.0 g  $\text{MgCl}_2$
55. a. 16% (v/v) ethanol  
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water  
b. 1.5M  $\text{KNO}_3$   
c. 0.100M  $\text{MgCl}_2$
58. The effective molality of the  $\text{Ca}(\text{NO}_3)_2$  solution is 3m. The effective molality of the  $\text{NaNO}_3$  solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g  $\text{H}_2\text{O}$  to 32.0 g  $\text{CH}_3\text{OH}$ .
63. a. 100.26°C  
b. 101.54°C
64. a.  $-4.46^\circ\text{C}$   
b.  $-2.2^\circ\text{C}$
65. a.  $-1.1^\circ\text{C}$   
b.  $-0.74^\circ\text{C}$   
c.  $-1.5^\circ\text{C}$

### Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

# Mastering Chemistry Answers Ch 23

**David R. Klein**



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**Mastering Organic Chemistry** Gerald W. Gibson, 1979 *Mastering Chemistry* Jason Chin, Vault Reports (Firm), Job Vault, 1998 The periodic table never looked so good From the properties of matter and mole calculations to the kinetics of reactions and enthalpy change this guide demystifies the complexities of chemistry through easy to follow charts graphs and sample problems High School Chemistry Unlocked The Princeton Review, 2016-10-18 UNLOCK THE SECRETS OF CHEMISTRY with THE PRINCETON REVIEW High School Chemistry Unlocked focuses on giving you a wide range of key lessons to help increase your understanding of chemistry With this book you ll move from foundational concepts to complicated real world applications building confidence as your skills improve End of chapter drills will help test your comprehension of each facet of chemistry from atoms to alpha radiation Don t feel locked out Everything You Need to Know About Chemistry Complex concepts explained in straightforward ways Walk throughs of sample problems for all topics Clear goals and self assessments to help you pinpoint areas for further review Guided examples of how to solve problems for common subjects Practice Your Way to Excellence 165 hands on practice questions seeded throughout the chapters and online Complete answer explanations to boost understanding Bonus online questions similar to those you ll find on the AP Chemistry Exam and the SAT Chemistry Subject Test High School Chemistry Unlocked covers Building blocks of matter Physical behavior of matter Chemical bonding Chemical reactions Stoichiometry Solutions Acids and bases Equilibrium Organic chemistry Radioactivity and more *Mastering Chemistry* Gerald W. Gibson, 1975 **Organic Chemistry** Marye Anne Fox, James K. Whitesell, 2004 Accompanying CD ROM has been enhanced with updated animated illustrations to accompany the presentations and Chem3D files for helpful structure visualization Page 4 of cover **Master the GED: Mastering the Science Test** Peterson's, 2010-08-01 Peterson s Master the GED Mastering the Science Test offers readers a complete look at the GED Science Test Readers will learn all about the GED Science test including What s tested and what s not tested Formats used Subject areas Question types based on the four skill areas Application questions Questions based on visual depictions General test taking strategies to score high Master the GED Mastering the Science Test is part of Master the GED 2011 which offers readers 3 full length practice tests and in depth subject review for each of the GED tests Language Arts Writing Parts I and II Language Arts Reading Social Studies including Canadian history and government Science and Mathematics Parts I and II as well as top test taking tips to score high on the GED **GED Test For Dummies** Murray Shukyn, Dale E. Shuttleworth, Achim K. Krull, 2014-10-06 Score higher on the GED with this book online practice If you re preparing for this all important exam GED Test For Dummies with Online Practice gets you up and running on everything you can expect on test day from overviews of the test sections to invaluable reviews and test taking strategies for all the subjects covered and everything in between In the book you ll find hands on digestible information for navigating your way through the Language Arts Reading and Writing Tests Social Studies Test Mathematics Test and Science Test Whether

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