

Lewis Structure Worksheet

1. Draw the Lewis Structure for each of the following elements.

| Element | Lewis Structure | Element | Lewis Structure | Element | Lewis Structure |
|-----------|-------------------|----------|------------------------|---------|-------------------------|
| Lithium | $\text{Li} \cdot$ | Nitrogen | $\cdot \text{N} \cdot$ | Silicon | $\cdot \text{Si} \cdot$ |
| Beryllium | Be | Fluorine | $\cdot \text{F} \cdot$ | Neon | $\cdot \text{Ne} \cdot$ |

2. Draw the Lewis Dot Diagram for each of the following elements. Determine how many electrons the atom has to gain or lose to become stable (fill or lose outer orbit to make 8). Determine the charge that the ion will have (recall: gain electrons = "-" charge, lose electrons = "+" charge).

| Element | Lewis Structure | Needed for Stability | Charge on Ion |
|-----------|------------------------|----------------------|----------------------|
| Sodium | $\text{Na} \cdot$ | Lose 1 | +1 |
| Oxygen | $\cdot \text{O} \cdot$ | Gain 2 | -2 |
| Aluminum | $\text{Al} \cdot$ | Lose 3 | +3 |
| Helium | He | N/A | N/A |
| Potassium | $\text{K} \cdot$ | Lose 1 | +1 |
| Carbon | $\cdot \text{C} \cdot$ | 4 | Does not form an ion |

Lewis Dot Structure Worksheet With Answers

Karin Nielsen-Saines

A red circular graphic with a gradient, appearing as a partial circle or a thick curved line, located to the right of the author's name.

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