

Board Question Paper: March 2014

Chemistry

Time: 3 Hours

Total Marks: 70

Note:

- All questions are compulsory.
- Answer to the two sections are to be written in the same answer book.
- Figure to the right hand side indicate full marks.
- Write balanced chemical equations and draw neat and labelled diagrams wherever necessary.
- Every new question must be started on a new page.
- Use of logarithmic table is allowed

SECTION – I

Q.1. Answer any ONE of the following:

[7]

- What is 'boiling point'?
Derive a relation between ΔH and ΔU for a chemical reaction.
Draw neat labelled diagram of calomel electrode.
Resistance and conductivity of a cell containing 0.001 M KCl solution at 298 K are 1500 Ω and 1.46×10^{-4} S. cm^{-1} respectively. What is the cell constant?
- Write molecularity of the following reaction:
 $2\text{NO}_{(\text{g})} + \text{O}_{2(\text{g})} \longrightarrow 2\text{NO}_{2(\text{g})}$
What is 'calcination'? How does it differ from 'roasting'?
Write resonating structures of ozone.
The decomposition of $\text{N}_2\text{O}_{5(\text{g})}$ at 320 K according to the following equation follows first order reaction:
$$\text{N}_2\text{O}_{5(\text{g})} \rightarrow 2\text{NO}_{2(\text{g})} + \frac{1}{2} \text{O}_{2(\text{g})}$$

The initial concentration of $\text{N}_2\text{O}_{5(\text{g})}$ is 1.24×10^{-2} mol. L^{-1} and after 60 minutes, 0.20×10^{-2} mol. L^{-1} . Calculate the rate constant of the reaction at 320 K.

Q.2. Answer any THREE of the following:

[9]

- One mole of a gas expands by 3 L against a constant pressure of 3 atmosphere. Calculate the work done in:
 - L. atmosphere
 - Joules
 - Calories
- Calculate the amount of CaCl_2 (van't Hoff factor $i = 2.47$) dissolved in 2.5 L solution so that its osmotic pressure at 300 K is 0.75 atmosphere.
Given: Molar mass of CaCl_2 is 111 g. mol^{-1} .
 $R = 0.082$ L. atm. $\text{K}^{-1} \text{mol}^{-1}$
- Describe anomalous behaviour of fluorine with the other elements of group 17 with reference to:
 - Hydrogen bonding
 - Oxidation state
 - Polyhalide ions

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Alexander Durie



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