STD XII - CHEMISTRY

Time: 3 hrs.l.

MARCH - 2013

[Marks: 70

7

SECTION - I

Q. 1.	Select and write the most appropriate answers from the given	
	alternatives for each sub-question :	
	In hade control which the state of the state	

In body centred cubic structure the space occupied is about (b) 53% (d) 32%

(a) 68% (c) 38% (ii) For a gaseous reaction the unit of rate of reaction is

(a) L atm s⁻¹ (b) atm $mol^{-1} s^{-1}$ $(c, stm s^{-1})$

(iii) Which of the following compounds contains S = O as well as S = S bonds?

(a) Sulphuric acid (b) Thiosulphuric aci (c) Sulphurous acid

(d) Thiosulphurous acid

(iv) Which of the following solutions shows maximum depression in freezing point? (a) 0.5 M Li₂SO₄ (b) 1 M NaCl (c) 0.5 M Al₂(SO₄)₃ (d) 0.5 M BaCl₂

(v) For a chemical reaction dS = -0.035 kJ/K and dH = -20 kJ. At what temperature does the reaction turn nonspontaneous? (a) 5.14 K (b) 57.14 K (d) 5714.0 K

(c) 571.4 K (vi) The standard e.m.f. of the following cell is 0.463 V

 $Cu_{(s)}$ $Cu_{(\alpha q)}^{++}$ $Ag_{(\alpha q)}^{++}$ $Ag_{(s)}$. What is the standard potential of Cu electrode?

(a) 1.137 V (b) 0.337 V (e) 0.463 V

(d) -0.463 V

(vii) Fe₂O₃ is reduced to spongy iron near the top of blast furnace by (a) H. (b) CaO (c) SiO₂ (d) CO

Q. 2. Answer any SIX of the following:

(i) Distinguish between crystalline solid and amorphous solid.

(ii) State Kohlrausch Law and write mathematical expression of molar conductivity of the given solution at infinite dilution.

(iii) Write cell reactions in lead storage battery during discharge.

(iv) Draw structures and write geometry of PCl₃ and PCl₅.

(v) Prove that ΔH = ΔU + Δ nRT. What is the condition under which ΔU = ΔH?

(vi) Mention names and formulae of two ores of aluminium.

(vii) Derive the relationship between relative lowering of vapour pressure and molar mass of nonvolatile solute.

(viii) What is pseudo first order reaction? Give one example of it.

Q. 3. Answer any THREE of the following:

[9]

[12]

(i) Calculate the mole fraction and molality of HNO3 in a solution containing 12.2% HNO3. (Given - atomic masses: H = 1, N = 14, O = 16)

(ii) Consider the reaction —

 $3I_{(ag)}^{-} + S_2O_8^{2-} \longrightarrow I_{3(ag)}^{-} + 2SO_{4(ag)}^{2-}$ at particular time t, $\frac{d[SO_4^{2-}]}{d} = 2.2 \times 10^{-2} \text{ M/s}.$

What are the values of the following at the same time?

(a)
$$-\frac{d[I]}{dt}$$

(b)
$$-\frac{d\left[S_2O_8^{2-}\right]}{dt}$$

(c)
$$-\frac{d\left[I_3^-\right]}{dt}$$

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- 1. Understanding the eBook Hsc Board Chemistry Paper Solution Of 2013
 - The Rise of Digital Reading Hsc Board Chemistry Paper Solution Of 2013
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Hsc Board Chemistry Paper Solution Of 2013
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - o Features to Look for in an Hsc Board Chemistry Paper Solution Of 2013
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Hsc Board Chemistry Paper Solution Of 2013
 - Personalized Recommendations
 - Hsc Board Chemistry Paper Solution Of 2013 User Reviews and Ratings
 - Hsc Board Chemistry Paper Solution Of 2013 and Bestseller Lists

- 5. Accessing Hsc Board Chemistry Paper Solution Of 2013 Free and Paid eBooks
 - Hsc Board Chemistry Paper Solution Of 2013 Public Domain eBooks
 - Hsc Board Chemistry Paper Solution Of 2013 eBook Subscription Services
 - Hsc Board Chemistry Paper Solution Of 2013 Budget-Friendly Options
- 6. Navigating Hsc Board Chemistry Paper Solution Of 2013 eBook Formats
 - o ePub, PDF, MOBI, and More
 - Hsc Board Chemistry Paper Solution Of 2013 Compatibility with Devices
 - Hsc Board Chemistry Paper Solution Of 2013 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Board Chemistry Paper Solution Of 2013
 - Highlighting and Note-Taking Hsc Board Chemistry Paper Solution Of 2013
 - Interactive Elements Hsc Board Chemistry Paper Solution Of 2013
- 8. Staying Engaged with Hsc Board Chemistry Paper Solution Of 2013
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Board Chemistry Paper Solution Of 2013
- 9. Balancing eBooks and Physical Books Hsc Board Chemistry Paper Solution Of 2013
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Board Chemistry Paper Solution Of 2013
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Board Chemistry Paper Solution Of 2013
 - Setting Reading Goals Hsc Board Chemistry Paper Solution Of 2013
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Board Chemistry Paper Solution Of 2013
 - Fact-Checking eBook Content of Hsc Board Chemistry Paper Solution Of 2013
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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