

Chapter 11: Gas Law Worksheet Answer Key

Complete the following calculation by list the given information, rewriting the formula to solve for the unknown, and plugging in the information (number and units), and writing the answer significantly.

1. A sample of 0.500 moles of gas is placed in a container of volume of 2.50 L. What is the pressure of the gas in torr if the gas is at 25 °C? (Ideal Gas Law)

$$P = \frac{nRT}{V} = \frac{(0.500 \text{ moles}) (0.0821 \text{ L} \cdot \text{atm}) (298 \text{ K})}{(\text{mole} \cdot \text{K}) (2.50 \text{ L})} = P = 4.89 \text{ atm} = 3720 \text{ torr}$$

2. If 2.00 mol of gas occupies 4.50 L at STP. How much of the same gas will occupy 3.00 L at STP? (Avogadro's Law)

$$n_2 = \frac{V_2 n_1}{V_1} = \frac{(3.00 \text{ L}) (2.00 \text{ moles})}{(4.50 \text{ L})} = 1.33 \text{ moles}$$

3. Determine the partial pressures of each of the gases in the following mixture: 17.04 g NH₃, 40.36 g Ne and 19.00 g F₂. The gases are at 1.50 atm of pressure. (Dalton's Law of Partial Pressure & Mole Ratios)

$$\frac{17.04 \text{ g NH}_3}{17.04 \text{ g NH}_3} \left| \frac{1 \text{ mole NH}_3}{17.04 \text{ g NH}_3} \right| \left| \frac{1.000 \text{ mole NH}_3}{3.500 \text{ moles}} \right| \quad \frac{1.000 \text{ mole NH}_3}{3.500 \text{ moles}} = 0.2857 \times 1.50 \text{ atm} = 0.429 \text{ atm NH}_3$$

$$\frac{40.36 \text{ g Ne}}{20.18 \text{ g Ne}} \left| \frac{1 \text{ mole Ne}}{20.18 \text{ g Ne}} \right| \left| \frac{2.000 \text{ moles Ne}}{3.500 \text{ moles}} \right| \quad \frac{2.000 \text{ moles Ne}}{3.500 \text{ moles}} = 0.5714 \times 1.50 \text{ atm} = 0.857 \text{ atm Ne}$$

$$\frac{19.00 \text{ g F}_2}{38.00 \text{ g F}_2} \left| \frac{1 \text{ mole F}_2}{38.00 \text{ g F}_2} \right| \left| \frac{0.5000 \text{ moles F}_2}{3.500 \text{ moles}} \right| \quad \frac{0.5000 \text{ moles F}_2}{3.500 \text{ moles}} = 0.1429 \times 1.50 \text{ atm} = 0.214 \text{ atm F}_2$$

$$1.000 \text{ moles} + 2.000 \text{ moles} + 0.5000 \text{ moles} = 3.500 \text{ moles}$$

4. If at 1.00 atm of pressure water boils at 100. °C, at what temperature would water boil if the pressure is 600. torr? (This shows why food doesn't cook well at higher elevations) (Gay-Lussac's Law)

$$T_2 = \frac{T_1 P_2}{P_1} = \frac{(373 \text{ K}) (0.789 \text{ atm})}{(1.00 \text{ atm})} = 294 \text{ K}$$

5. Calculate the volume of 40.6 g of F₂ at STP. (Ideal Gas Law)

$$\frac{40.6 \text{ g F}_2}{38.00 \text{ g F}_2} \left| \frac{1 \text{ mole F}_2}{38.00 \text{ g F}_2} \right| \left| \frac{22.4 \text{ L F}_2}{1 \text{ mole F}_2} \right| \quad \frac{22.4 \text{ L F}_2}{1 \text{ mole F}_2} = 23.9 \text{ L F}_2$$

or

$$\frac{40.6 \text{ g F}_2}{38.00 \text{ g F}_2} \left| \frac{1 \text{ mole F}_2}{38.00 \text{ g F}_2} \right| \left| \frac{1.07 \text{ mole F}_2}{1.07 \text{ mole F}_2} \right| \quad \frac{1.07 \text{ mole F}_2}{1.07 \text{ mole F}_2} = 24.0 \text{ L F}_2$$

$$V = \frac{nRT}{P} = \frac{(1.07 \text{ moles}) (0.0821 \text{ L} \cdot \text{atm}) (273 \text{ K})}{(\text{mole} \cdot \text{K}) (1.00 \text{ atm})} = 24.0 \text{ L F}_2$$

6. If the winter pressure that most car tire tires should be at to wear evenly is 32.0 psi, what is the pressure in atm? (Pressure Conversion)

$$\frac{32.0 \text{ psi}}{14.7 \text{ psi}} \left| \frac{1.00 \text{ atm}}{14.7 \text{ psi}} \right| \quad \frac{1.00 \text{ atm}}{14.7 \text{ psi}} = 2.18 \text{ atm}$$

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