

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Goals: (1) (10-11 P. 41) Explain a new scientific theory of atomic-scale phenomena (atomic theory) by describing changes in the atomic model over time, including those changes that are consistent with experimental evidence. (10-11 P. 42) Explain the scientific theory of atomic-scale phenomena, including the scientific theories in terms of particles, motions, interactions, and the evidence supporting those particles and motions. (10-11 P. 43) Explain changes and motions within the atom. (10-11 P. 44) Explain properties of atoms and their motions within particles, including the composition of their components. (10-11 P. 45) Explain the scientific evidence of the four fundamental forces (gravitational, electromagnetic, weak nuclear, strong nuclear).

ANSWER KEY

Directions: Match the atomic theory selected on the correct cardstock.

Modeler	Scientist(s)	Atomic Theory's Basic Model
1C	Democritus (460 BC)	A. Everything is made of atoms that cannot be created, divided, or destroyed.
4A	Boyle (1627)	B. Matter is made of atoms, which are mostly empty space, and are separated by constant intermolecular attractive forces.
10B	Thomson (1897)	C. "Atoms are indivisible particles."
7	Rutherford (1911)	D. Group Number = Valence Electrons
10A	Bohr (1913)	E. Electrons move
10C	Schrodinger (1926)	F. Atoms are mostly empty space, and the subatomic particles move.

Directions: Use the periodic table to answer the following questions:

- 10C** = 11-18. The number of protons equals the number of electrons.
- Group 18 has the atomic number all equal that **Number of electrons**.
- 10C** = **10-18** protons are positive (positively charged) than most of which other **electrons** are negative.

Energy And Electrons Study Guide Answers

Roman Wölfel



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