Unit 3 Energy Changes and Rates of Reaction

Solutions to Practice Problems in Chapter 5 Energy Changes

Calculating the Absorption of Heat (Student textbook page 281)

 How much heat must be added to a 36.2 g sample of iron to increase its temperature by 250.0°C?

What Is Required?

You have to calculate the amount of heat that must be added to a sample of iron to increase its temperature by 250.0°C.

What Is Given?

You know the mass of iron: m = 36.2 g

You know the change in the temperature: $\Delta T = 250.0^{\circ}$ C

You have the specific heat capacity of iron from Table 5.1 on page 280 of the student text: $c_{Fe} = 0.449 \text{ J/g} \cdot ^{\circ}\text{C}$

Plan Your Strategy	Act on Your Strategy
Substitute the given information into the formula $Q = mc\Delta T$. Solve for Q .	$Q = mc\Delta T$ = $(36.2 \text{ f}) \left(0.449 \frac{\text{J}}{\text{f} \cdot 26}\right) (250.026)$ = 4063.45 J = 4.06 kJ The iron must gain 4060 J or 4.06 kJ of heat energy.

Check Your Solution

Check that the information has been substituted correctly. The answer has the correct number of significant digits.

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