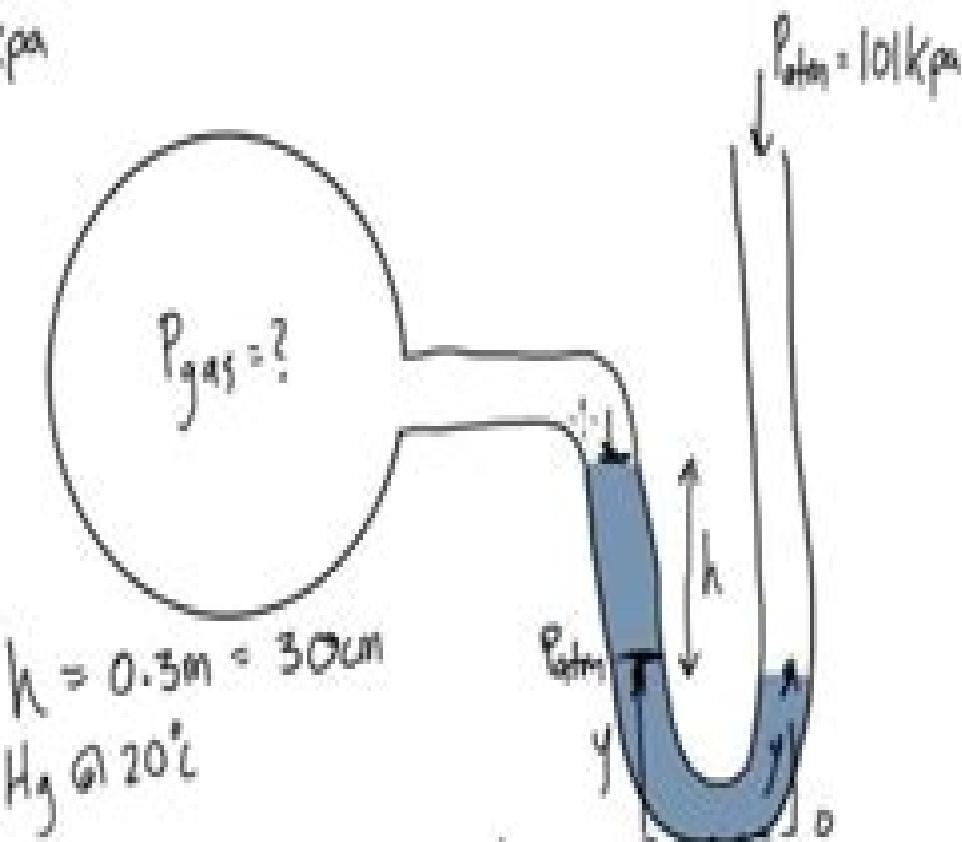


$$\begin{aligned}
 P_{gas} &= P_{atm} + \rho gh \\
 &= 101 \text{ kPa} + (13,550 \frac{\text{N}}{\text{m}^2}) \left(9.81 \frac{\text{m}}{\text{s}^2} \right) (0.3) \\
 &= 101 \text{ kPa} + 39.877 \text{ kPa}
 \end{aligned}$$

$$P_{gas} = 141 \text{ kPa}$$



$$\begin{aligned}
 h &= 0.3 \text{ m} = 30 \text{ cm} \\
 \text{Hg} @ 20^\circ\text{C}
 \end{aligned}$$

$$\begin{aligned}
 P_{gas} &= P_{atm} - \rho gh \\
 &= 101 \text{ kPa} - (13,550)(9.81)(0.3) \\
 &= 101 - 39.877 \\
 P_{gas} &= 61 \text{ kPa}
 \end{aligned}$$

Manometer Example Problems

**R. Byron Bird, Warren E. Stewart, Edwin
N. Lightfoot**

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