

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Goals: (a) (a.i) P. 4.1 - Explain a new scientific theory of atomic-scale phenomena (atomic theory) by describing changes in the atomic model over time, including those changes that are consistent with experimental evidence. (a.i.ii) P. 4.2 - Explain the scientific theory of atomic-scale phenomena, including the scientific changes in terms of changing scientific understandings, scientific evidence, and the role of theory in science. (a.i.iii) P. 4.3 - Explain changes in the atomic model over time, including the scientific changes in terms of changing scientific understandings, scientific evidence, and the role of theory in science. (a.i.iv) P. 4.4 - Explain the scientific theory of atomic-scale phenomena, including the scientific changes in terms of changing scientific understandings, scientific evidence, and the role of theory in science.

ANSWER KEY

Directions: Match the atomic theory selected on the correct understanding.

Model	Scientist	Atomic Theory
1	Democritus (460 BC)	A. Everything is made of atoms that cannot be created, divided, or destroyed.
2	Boyle (1627)	B. Matter is made of atoms that are constantly moving and are separated by spaces.
3	Newton (1642)	C. Matter is made of particles (molecules).
4	Thomson (1857)	D. Matter is made of particles (molecules).
5	Rutherford (1871)	E. Matter is made of particles (molecules).
6	Boyle (1627)	F. Matter is made of particles (molecules).
7	Thomson (1857)	G. Matter is made of particles (molecules).

Directions: Use the periodic table to answer the following questions:

- 100** = P. 4.1.11 - The number of protons equals the number of electrons.
- 100** = P. 4.1.11 - The number of protons equals the number of electrons.
- 100** = P. 4.1.11 - The number of protons equals the number of electrons.

Electrons In Atoms Answer Key Study Guide

Barbara Herlihy



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