

# BOARD QUESTION PAPER: MARCH 2022

## CHEMISTRY

Time: 3 Hrs

Max. Marks: 70

### General Instructions:

The question paper is divided into four sections.

- (1) **Section A:** Q. No. 1 contains Ten multiple choice type of questions carrying One mark each.  
Q. No. 2 contains Eight very short answer type of questions carrying One mark each.
- (2) **Section B:** Q. No. 3 to Q. No. 14 are Twelve short answer type of questions carrying Two marks each. (Attempt any Eight).
- (3) **Section C:** Q. No. 15 to Q. No. 26 are Twelve short answer type of questions carrying Three marks each. (Attempt any Eight).
- (4) **Section D:** Q. No. 27 to Q. No. 31 are Five long answer type of questions carrying Four marks each. (Attempt any Three).
- (5) Use of log table is allowed. Use of calculator is not allowed.
- (6) Figures to the right indicate full marks.
- (7) For each multiple choice type of question, it is mandatory to write the correct answer along with its alphabet, e.g. (a)...../(b)...../(c)...../(d)..... etc.

No marks(s) shall be given, if ONLY the correct answer or the alphabet of the correct answer is written. Only the first attempt will be considered for evaluation.

### SECTION – A

Q.1. Select and write the correct answer for the following multiple choice type of questions:

[10]

- (i) The co-ordination number of atoms in body centred cubic structure (bcc) is \_\_\_\_\_.  
(a) 4                      (b) 6                      (c) 8                      (d) 12
- (ii) In calculating osmotic pressure, the concentration of solute is expressed in \_\_\_\_\_.  
(a) molarity                      (b) molality  
(c) mole fraction                      (d) percentage mass
- (iii) The enthalpy change for the chemical reaction  $\text{H}_2\text{O}_{(s)} \longrightarrow \text{H}_2\text{O}_{(g)}$  is called enthalpy of \_\_\_\_\_.  
(a) vapourisation                      (b) fusion  
(c) combustion                      (d) sublimation
- (iv) Which of the following transition element shows maximum oxidation state?  
(a) Sc                      (b) Fe                      (c) Mn                      (d) V
- (v) The correct formula for the complex compound, sodium hexacyanoferrate (III) is \_\_\_\_\_.  
(a)  $\text{Na} [\text{Fe}(\text{CN})_6]$                       (b)  $\text{Na}_2 [\text{Fe}(\text{CN})_6]$   
(c)  $\text{Na}_3 [\text{Fe}(\text{CN})_6]$                       (d)  $\text{Na}_4 [\text{Fe}(\text{CN})_6]$

# Hsc Board Paper

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## **Table of Contents Hsc Board Paper**

1. Understanding the eBook Hsc Board Paper
  - The Rise of Digital Reading Hsc Board Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Board Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc Board Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Board Paper
  - Personalized Recommendations

- Hsc Board Paper User Reviews and Ratings
- Hsc Board Paper and Bestseller Lists
- 5. Accessing Hsc Board Paper Free and Paid eBooks
  - Hsc Board Paper Public Domain eBooks
  - Hsc Board Paper eBook Subscription Services
  - Hsc Board Paper Budget-Friendly Options
- 6. Navigating Hsc Board Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - Hsc Board Paper Compatibility with Devices
  - Hsc Board Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Hsc Board Paper
  - Highlighting and Note-Taking Hsc Board Paper
  - Interactive Elements Hsc Board Paper
- 8. Staying Engaged with Hsc Board Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hsc Board Paper
- 9. Balancing eBooks and Physical Books Hsc Board Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Hsc Board Paper
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Board Paper
  - Setting Reading Goals Hsc Board Paper
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Board Paper
  - Fact-Checking eBook Content of Hsc Board Paper

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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