

1. According to the Brønsted—Lowry definition, which chemical species can function *both* as an acid and as a base?

- $\text{Cl}^-$
- $\text{SO}_4^{2-}$
- $\text{NH}_4^+$
- $\text{HCO}_3^-$
- $\text{H}_3\text{O}^+$

2. In the reaction  $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$  which is an acid—base conjugate pair?

- $\text{H}_2\text{O}$  and  $\text{HCN}$
- $\text{H}_2\text{O}$  and  $\text{OH}^-$
- $\text{CN}^-$  and  $\text{H}_2\text{O}$
- $\text{HCN}$  and  $\text{OH}^-$

3. Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX

- It is less basic than a 1 M solution of NaY.
- It is more basic than a 1 M solution of NaY.
- It yields a neutral solution.
- It is more concentrated than a 1 M solution of NaY.

4. The pH of a 0.03 M HCl solution is

|    |      |
|----|------|
| a. | 1.5  |
| b. | 2.5  |
| c. | 3.5  |
| d. | 12.5 |

5. The pH of a  $1.0 \times 10^{-3}$  M  $\text{Ba}(\text{OH})_2$  solution at  $25^\circ\text{C}$  is

|    |      |
|----|------|
| a. | 2.7  |
| b. | 3.0  |
| c. | 11.0 |
| d. | 11.3 |

6. Which is the strongest acid?

- $\text{HClO}$
- $\text{HClO}_2$
- $\text{HClO}_3$
- $\text{HClO}_4$

7. Which salt reacts with water (hydrolyzes) to produce a basic solution?

a.  $\text{NaC}_2\text{H}_3\text{O}_2$   
b.  $\text{NaNO}_3$   
c.  $\text{NH}_4\text{Cl}$   
d.  $\text{BaSO}_4$

8. The weakest of the bases listed is

| Acid                          | Conjugate Base                | K <sub>a</sub> (Ionization Constant of Acid) |
|-------------------------------|-------------------------------|----------------------------------------------|
| HCl                           | Cl <sup>-</sup>               | 100% ionized                                 |
| HSO <sub>4</sub> <sup>-</sup> | SO <sub>4</sub> <sup>2-</sup> | 1.2 x 10 <sup>-2</sup>                       |
| H <sub>2</sub> S              | HS <sup>-</sup>               | 5.7 x 10 <sup>-8</sup>                       |
| HS <sup>-</sup>               | S <sup>2-</sup>               | 1.2 x 10 <sup>-14</sup>                      |

- $\text{Cl}^-$
- $\text{CN}^-$
- $\text{HS}^-$
- $\text{S}^{2-}$
- $\text{SO}_3^{2-}$

| Number of hauls | <i>M. undulatus</i> (%) | <i>M. setulosus</i> (%) | <i>M. setulosus</i> + <i>M. undulatus</i> (%) |
|-----------------|-------------------------|-------------------------|-----------------------------------------------|
| 1               | 100                     | 0                       | 100                                           |
| 2               | 80                      | 20                      | 100                                           |
| 3               | 40                      | 60                      | 100                                           |
| 4               | 10                      | 90                      | 100                                           |
| 5               | 5                       | 95                      | 100                                           |
| 6               | 2                       | 98                      | 100                                           |
| 7               | 1                       | 99                      | 100                                           |
| 8               | 0                       | 100                     | 100                                           |
| 9               | 0                       | 100                     | 100                                           |
| 10              | 0                       | 100                     | 100                                           |

10. Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate  $\text{KC}_2\text{H}_3\text{O}_2$  is less alkaline than a 0.10 M solution of potassium cyanide, KCN?

- Hydrocyanic acid is a weaker acid than acetic acid.
- Hydrocyanic acid is less soluble in water than acetic acid.
- Cyanides are less soluble than acetates.
- Acetic acid is a weaker acid than hydrocyanic acid.

11. In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of

- 0.100 M sodium solution,
- 0.0500 M sodium hydroxide solution,
- 0.0500 M benzoic acid solution,
- 0.0500 M sodium benzoate solution.

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**M Woodhall**



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